TOPIC: **DENSITY**

1. Calculate the **density** of each of the following:
   a. 250 cm^3 of a solution with a mass of 500 g

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}} = \frac{500\ g}{250\ cm^3} = 2\ g/cm^3$$

   b. A 6.75 g solid with a volume of 5 cm^3

$$= _____ = \frac{_____\ g}{cm^3} = _____\ g/cm^3$$

   c. 50.0 g of a gas which occupies a volume of 8 cm^3

$$= _____ = \frac{_____\ g}{cm^3} = _____\ g/cm^3$$

2. Calculate the **volume** of each of the following:
   a. 26.5 g of a solution with a density of 2 g/cm^3

$$= _____ = \frac{_____\ g}{g/cm^3} = _____\ cm^3$$

   b. A 3.40 g solid with a density of 10 g/cm^3

$$= _____ = \frac{_____\ g}{g/cm^3} = _____\ cm^3$$

3. Calculate the **mass** of each of the following:
   a. A solid with a volume of 1.68 cm^3 and a density of 9.2 g/cm^3

$$= _____ \text{ x } _____$$

$$= _____\ g/cm^3 \text{ x } _____\ cm^3$$

$$= _____\ g$$

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b. An 80 cm^3 of a solution with a density of 5.80 g/cm^3

= _____ x _____

= _____ g/cm^3 x _____ cm^3

= _____ g

a. A solid with a density of 4 g/cm^3 and dimensions of 2.5 cm x 2.5 cm x 2.5 cm

= _____ x _____

= _____ g/cm^3 x _____ cm^3

= _____ g

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