

Name: _____ Date: _____ Class: _____

How Many Particles?

Directions: Each of the following compounds is soluble in water. Determine the number of moles of particles (ions or molecules) present per mole of the compound being dissolved in water.

1) NaCl _____

2) BaCl₂ _____

3) C₆H₁₂O₆ _____

4) LiF _____

5) NH₃ _____

6) KCl _____

7) Ca(OH)₂ _____

8) MgBr₂ _____

9) LiOH _____

10) C₂H₆O₂ _____

11) AgNO₃ _____

12) RbC₂H₃O₂ _____

13) CH₃OH _____

14) NH₄CO₃ _____

15) K₃PO₄ _____

16) Na₂CO₃ _____

17) C₆H₁₄O₇ _____

18) Ba(NO₃)₂ _____

19) Mg(ClO₃)₂ _____

20) Ba(C₂H₃O₂)₂ _____



IF YOU'RE NOT
A PART OF THE
solution
YOU MUST BE A
PART OF THE
precipitate.