

Balancing Equations Practice

What are the steps for balancing equations?

1. List the _____ found on both sides of the equation. List them in a column and in the _____ order.
2. Count the number of _____ of each element and write them.
3. Add _____ and multiply these by the subscripts. Record the new numbers of atoms. And continue until the equation is balanced.

Balance the following equations by adding the correct coefficients. (NOTE: Type 1 if we would normally not write a coefficient.) Do the work on a separate sheet of paper.

$\underline{\quad} \text{Fe} + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{Fe}_2\text{O}_3$	$\underline{\quad} \text{Fe} + \underline{\quad} \text{Cl}_2 \rightarrow \underline{\quad} \text{FeCl}_3$
$\underline{\quad} \text{P} + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{P}_2\text{O}_5$	$\underline{\quad} \text{Ag}_2\text{O} \rightarrow \underline{\quad} \text{Ag} + \underline{\quad} \text{O}_2$
$\underline{\quad} \text{S}_8 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{SO}_3$	$\underline{\quad} \text{K} + \underline{\quad} \text{H}_2\text{SO}_4 \rightarrow \underline{\quad} \text{K}_2\text{SO}_4 + \underline{\quad} \text{H}_2$