_____ 1. The mass of a molecule is found by
   a. adding the masses of all the atoms in the molecule.
   b. dividing the total mass of all the atoms in the molecule
      by 6.022×10^{23}.
   c. multiplying the total mass of all the atoms in the molecule
      by 6.022×10^{23}.
   d. dividing the total mass of all the atoms in the molecule by the total number of
      atoms.

_____ 2. For any molecule, formula unit, or ion, the sum of the average atomic masses of
   all the atoms represented in a formula is the
   a. formula mass.
   b. ionic mass.
   c. molecular mass.
   d. atomic mass.

_____ 3. What is the correct formula mass for zinc hydroxide, $Zn(OH)_2$?
   a. 83.41 g
   b. 99.41 g
   c. 83.41 amu
   d. 99.41 amu

_____ 4. One mole of potassium atoms has a mass of
   a. 1 g.
   b. 39.10 g.
   c. 78.20 g.
   d. 6.022×10^{23} g.

_____ 5. A compound's molar mass is numerically equal to
   a. the total number of atoms in the molecule.
   b. the total number of moles of the compound.
   c. its mass number.
   d. its formula mass.