

**DIRECTIONS: FOR EACH PROBLEM. DRAG THE VALUES INTO THE CORRECT PLACES.
THEN, CALCULATE THE ANSWER (round to 3 decimal places)**

1. How many grams are in 6.3 moles of sulfur? *32.07 g* *6.3 moles* *1 mole* *6.02×10^{23} atoms*

<div style="border: 1px solid black; width: 150px; height: 30px; margin: 0 auto;"></div>	<div style="border: 1px solid black; width: 150px; height: 30px; margin: 0 auto;"></div>	x	<div style="border: 1px solid black; width: 150px; height: 30px; margin: 0 auto;"></div>	=	<div style="border: 1px solid red; width: 150px; height: 30px; margin: 0 auto;"></div>
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2. How many moles are in 7 g of carbon? *7 g* *1 mole* *12.01 g* *6.02×10^{23} atoms*

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3. How many atoms are in 7.1 g of sodium? *6.02×10^{23} atoms* *7.1 g* *1mole* *22.99 g*

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4. How many atoms are in 0.35 moles of aluminum? *6.02×10^{23} atoms* *1 mole* *0.35 moles* *26.98 g*

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5. How many moles are in 22,000,000 atoms of calcium? *22,000,000 atoms* *6.02×10^{23} atoms* *1 mole* *40.08 g*

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