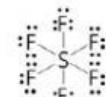
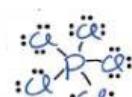
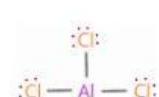
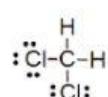
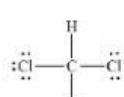
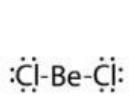
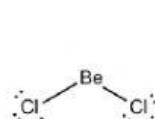
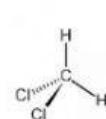
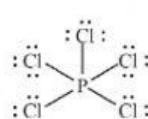
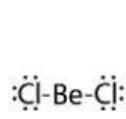
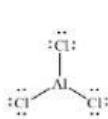
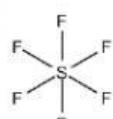
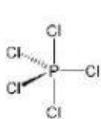


Question 2 (Intermolecular forces)

Complete the table with related answers. Name the type Intermolecular forces (IMF) that exist between molecules for each of the following species:

Formula	Covalent compounds				
	BeCl_2	AlCl_3	CH_2Cl_2	PCl_5	SF_6
Lewis structure					
Shape of molecule					
State the shape of molecule					
Net dipole moment, μ					
Polar/Non-polar					
State all the IMF between the molecules					

Answer: Drag the correct answer to the suitable column



Note:

Hydrogen bond = HB

Dipole-dipole force= DDF (for polar compound, μ not equal to 0)

London Dispersion forces = LDF (for non-polar compound, μ equal to 0)