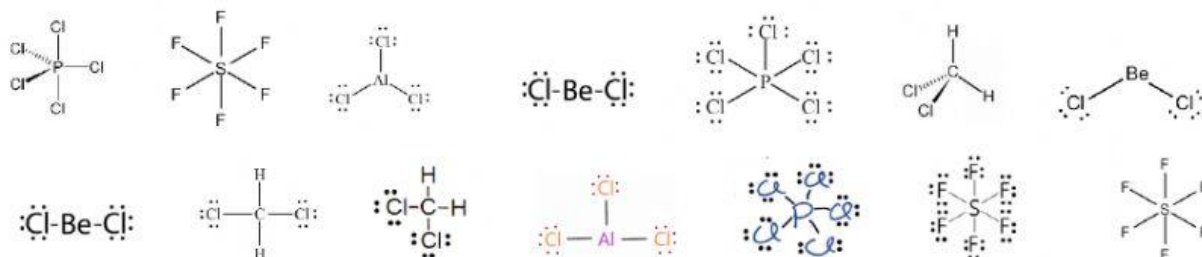


## Question 2 (Intermolecular forces)

Complete the table with related answers. Name the type Intermolecular forces (IMF) that exist between molecules for each of the following species:

Formula	Covalent compounds				
	<b>BeCl<sub>2</sub></b>	<b>AlCl<sub>3</sub></b>	<b>CH<sub>2</sub>Cl<sub>2</sub></b>	<b>PCl<sub>5</sub></b>	<b>SF<sub>6</sub></b>
Lewis structure					
Shape of molecule					
State the shape of molecule					
Net dipole moment, $\mu$					
Polar/Non-polar					
State <b>all the IMF</b> between the molecules					

**Answer:** Drag the correct answer to the suitable column



**Note:**

Hydrogen bond = HB

Dipole-dipole force= DDF (for polar compound,  $\mu$  not equal to 0)

London Dispersion forces = LDF (for non-polar compound,  $\mu$  equal to 0)