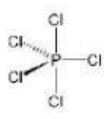
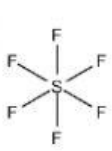
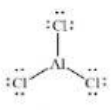
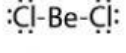
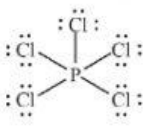
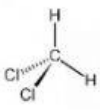
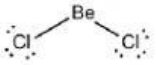
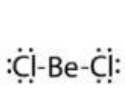
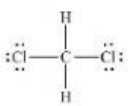
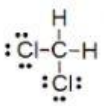
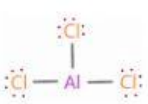
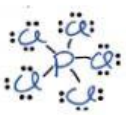
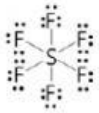
Question 2 (Intermolecular forces)

Complete the table with related answers. Name the type Intermolecular forces (IMF) that exist between molecules for each of the following species:

Formula	Covalent compounds				
	$BeCl_2$	**AlCl3**	**CH_2Cl_2**	**PCl_5**	**SF_6**
Lewis structure					
Shape of molecule					
State the shape of molecule					
Net dipole moment, μ					
Polar/Non-polar					
State all the IMF between the molecules					

Answer: Drag the correct answer to the suitable column


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Note:

Hydrogen bond = HB

Dipole-dipole force= DDF (for polar compound, $\mu \; not \; equal \; to \; 0$)

London Dispersion forces = LDF (for non-polar compound, $\mu \; equal \; to \; 0$)

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