

Average atomic mass

$$\text{Average atomic mass} = \frac{\% \text{ of abundance}}{100} \times \text{atomic mass}$$

Isotope	% abundance	mass number
	98.89%	12
	1.11%	13

Average Atomic
mass

$$\begin{aligned} &= \frac{98.89}{100} \times 12 + \\ &\quad \frac{1.11}{100} \times 13 \\ &= 11.8668 + 0.1443 \\ &= 12.0111 \text{ amu} \end{aligned}$$

Calculate the average atomic mass of below isotopes

Isotope	% Abundance	Isotope Mass
Oxygen-16	99.76%	16
Oxygen-18	0.200%	18

average atomic
mass

$$\begin{aligned}
 &= \frac{\boxed{}}{\boxed{}} \times \boxed{} + \\
 &\quad \frac{\boxed{}}{\boxed{}} \times \boxed{}
 \end{aligned}$$

Average Atomic
mass

$$\begin{aligned}
 &= \boxed{} + \boxed{} \\
 &= \boxed{} \text{ amu}
 \end{aligned}$$

Magnesium (Mg) Isotope	Percent Found in Nature
Mg-24	78.9%
Mg-25	10.0%

Average Atomic mass

$$\begin{aligned}
 &= \frac{\boxed{}}{\boxed{}} \times \boxed{} + \\
 &\quad \frac{\boxed{}}{\boxed{}} \times \boxed{} \\
 &= \boxed{} + \boxed{} \\
 &= \boxed{} \text{ amu}
 \end{aligned}$$

Isotopes	% Abundance	Mass in amu
Ga-69	60.11%	69
Ga-71	39.89%	71

Average Atomic mass

$$\begin{aligned}
 &= \frac{\boxed{}}{\boxed{}} \times \boxed{} + \\
 &\quad \frac{\boxed{}}{\boxed{}} \times \boxed{} \\
 &= \boxed{} + \boxed{} \\
 &= \boxed{} \text{ amu}
 \end{aligned}$$

Isotopes	% Abundance	Mass in amu
Copper-63	69.1%	62.9296
Copper-65	30.83%	64.9278

Average Atomic
mass

$$\begin{aligned}
 &= \frac{\boxed{}}{\boxed{}} \times \boxed{} + \\
 &\quad \frac{\boxed{}}{\boxed{}} \times \boxed{} \\
 &= \boxed{} + \boxed{} \\
 &= \boxed{} \text{ amu}
 \end{aligned}$$