

Name of the student:

Teacher: Sridhar Sriram

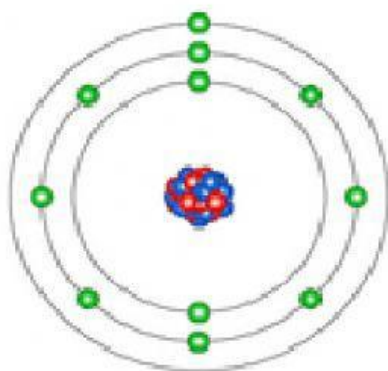
Class/Section:

Subject: Science (Chemistry)

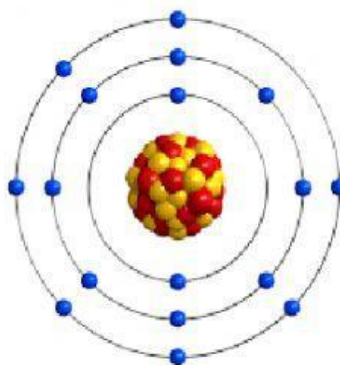
Chemical bonding

1. Lesson Story: A _____ is a lasting attraction between atoms, ions or molecules that enables the formation of chemical compounds. The bond may result from the electrostatic force of attraction between oppositely charged ions as in _____ or through the sharing of electrons as in _____. The strength of chemical bonds varies considerably; there are "strong bonds" or "primary bonds" such as covalent, ionic and metallic bonds, and "weak bonds" or "secondary bonds" such as dipole–dipole interactions, the London dispersion force and hydrogen bonding.

2. Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms.



Element



Element

Symbol

Symbol

Atomic number

Atomic number

Total electrons

Total electrons

Number of shells

Number of shells

Metal/Nonmetal

Metal/Nonmetal

Valence electrons

Valence electrons

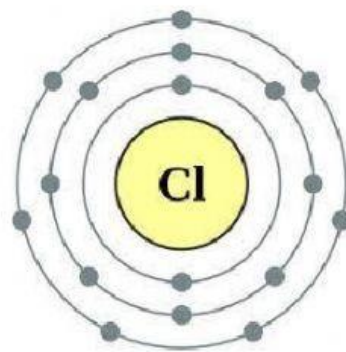
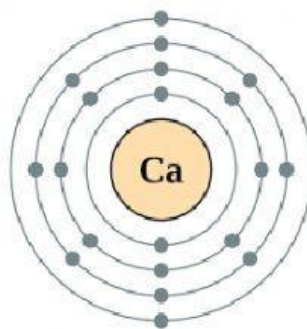
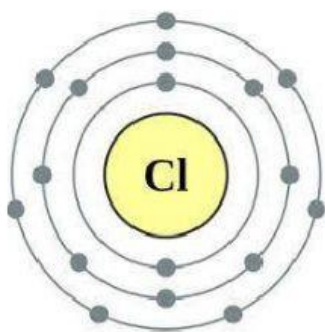
Lose /Gain electrons

Lose / Gain electrons

Type of bond

Type of bond

3. Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms



Element

Element

Symbol

Symbol

Atomic number

Atomic number

Total electrons

Total electrons

Number of shells

Number of shells

Metal/Nonmetal

Metal/Nonmetal

Valence electrons

Valence electrons

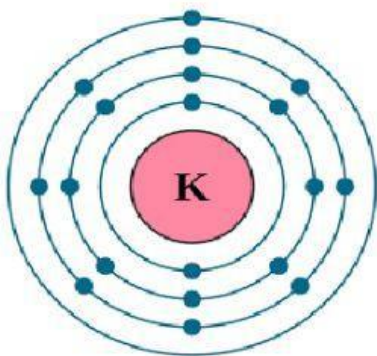
Lose /Gain

Lose / Gain

Type of bond

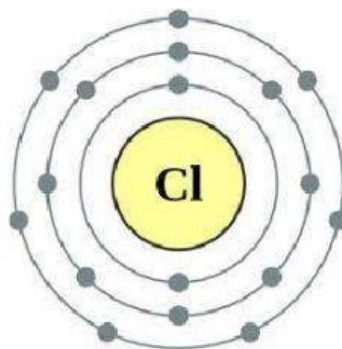
Type of bond

4. Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms



Element

Symbol



Element

Symbol

Atomic number

Total electrons

Number of shells

Metal/Nonmetal

Valence electrons

Lose /Gain

Type of bond

Atomic number

Total electrons

Number of shells

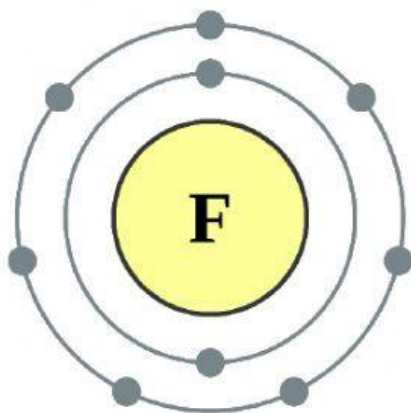
Metal/Nonmetal

Valence electrons

Lose / Gain

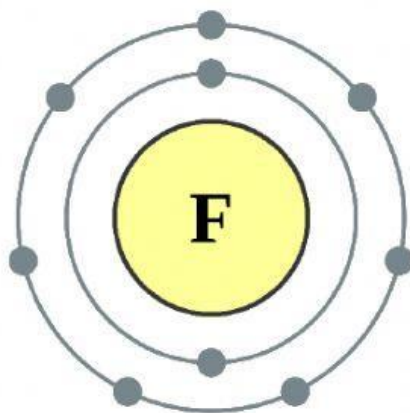
Type of bond

5. . Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms



Element

Symbol



Element

Symbol

Atomic number

Total electrons

Number of shells

Metal/Nonmetal

Valence electrons

Lose /Gain/Share

Type of bond

Atomic number

Total electrons

Number of shells

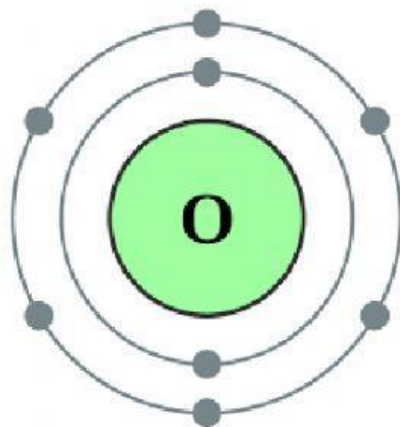
Metal/Nonmetal

Valence electrons

Lose / Gain/Share

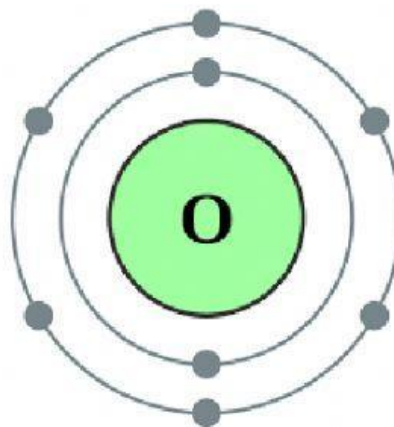
Type of bond

6. Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms



Element

Symbol



Element

Symbol

Atomic number

Atomic number

Total electrons

Total electrons

Number of shells

Number of shells

Metal/Nonmetal

Metal/Nonmetal

Valence electrons

Valence electrons

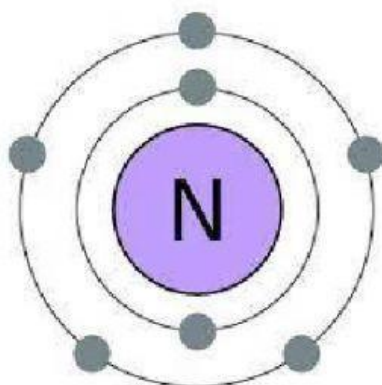
Lose /Gain/Share

Lose / Gain/Share

Type of bond

Type of bond

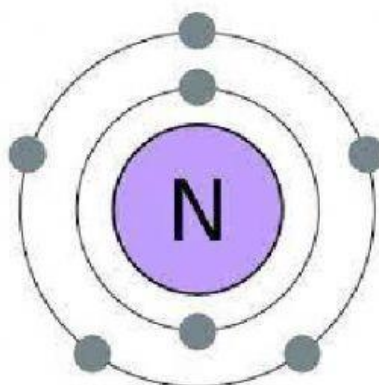
7. Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms



Element

Symbol

Atomic number



Element

Symbol

Atomic number

Total electrons

Total electrons

Number of shells

Number of shells

Metal/Nonmetal

Metal/Nonmetal

Valence electrons

Valence electrons

Lose /Gain/Share

Lose / Gain/Share

Type of bond

Type of bond