

1+						0
1						8
1	2+	Oxidation Number				8
1	2-	Valence Electrons				4
1	2	Family				He
1						4.003
3	4					10
Li	Be					Ne
Lithium	Beryllium					20.180
6.941	9.0122					
11	12					18
Na	Mg					Ar
Sodium	Magnesium					39.948
22.990	24.305					
19	20					36
K	Ca					Kr
Potassium	Calcium					83.80
39.098	40.078					

	3+	4+/-	3-	2-	1-	0
Oxidation Number						
Valence Electrons	3	4	5	6	7	8
Family	3	4	5	6	7	8

5	6	7	8	9	10
B	C	N	O	F	Ne
Boron	Carbon	Nitrogen	Oxygen	Fluorine	Neon
10.81	12.011	14.007	15.999	18.998	20.180
13	14	15	16	17	18
Al	Si	P	S	Cl	Ar
Aluminum	Silicon	Phosphorus	Sulfur	Chlorine	Argon
26.98	28.086	30.974	32.06	35.453	39.948
31	32	33	34	35	36
Ga	Ge	As	Se	Br	Kr
Gallium	Germanium	Arsenic	Selenium	Bromine	Krypton
69.72	72.61	74.922	78.96	79.904	83.80

Some Polyatomic Ions and their Oxidation Numbers					
1+	1-	2-	3-		
ammonium (NH ₄)	acetate (C ₂ H ₃ O ₂)	carbonate (CO ₃)	phosphate (PO ₄)		
	chlorate (ClO ₃)	sulfate (SO ₄)			
	hydroxide (OH)				
	nitrate (NO ₃)				
	bicarbonate (HNO ₃)				

	Symbols & Oxidation #	Formula	Name of Compound
15. Magnesium and chlorine			
16. Potassium and acetate			
17. Aluminum and sulfate			
18. Lithium and nitrogen			

19. An ion is an atom or group of atoms that has become electrically _____

20. When an atom loses an electron its charge is (*positive or negative*)

21. An ionic bond is the attraction between (*opposites, positive, neutral, or negative*) ions.

22. Ionic compounds are electrically (*charged, positive, neutral, or negative*).

23. The sum of the charges for an ionic compound is _____.

The two answers must be in the right order.

24. An ionic compound is the result of the bonding of a (*non-metal, metalloid, metal, noble gas*) with a (*non-metal, metalloid, metal*).