

EXERCISE PRACTICAL LAB REPORT

Part A: Preparation of diluted oxalic acid, $\text{H}_2\text{C}_2\text{O}_4$ solution

(Drag and Drop the following procedure to appropriate box in the correct order.)

Add distilled water to the volumetric flask up to the calibrated mark.
Rinse a 10 mL pipette with the $\text{H}_2\text{C}_2\text{O}_4$ stock solution.
Stir the solution with a glass rod and transfer it into a 250 mL volumetric flask.
Fill approximately 100 mL of distilled water into 250 mL beaker.
Pipette 10 mL of the $\text{H}_2\text{C}_2\text{O}_4$ stock solution and transfer it into the beaker containing the distilled water.
The diluted $\text{H}_2\text{C}_2\text{O}_4$ is ready for titration.
Rinse a 10 mL pipette with distilled water.
Stopper and shake the volumetric flask to obtain a homogenous solution.
Rinse the beaker and pour the content into the volumetric flask.

Step 1
Step 2
Step 3
Step 4
Step 5
Step 6
Step 7
Step 8
Step 9

PRACTICAL EXERCISE

250 mL of **diluted** Oxalic acid ($\text{H}_2\text{C}_2\text{O}_4$) is prepared by using 10 mL of $\text{H}_2\text{C}_2\text{O}_4$ **stock solution** in 250mL volumetric flask. 25 mL of **diluted** $\text{H}_2\text{C}_2\text{O}_4$ solution is used to titrate with 0.2 M NaOH solution. Volume of NaOH used in the titration is as follows:

(Fill in the empty column with correct answer)

Burette Reading (mL)	Gross	I	II	III
Final	41.70	25.30	30.70	27.80
Initial	16.00	00.00	5.50	2.50
Volume titrant (NaOH) used				

Calculation:

1. Average volume of NaOH used: mL

2. Chemical equation involved: *(Drag and Drop items to appropriate box)*



Drag the answer from here.

H_2O

NaOH

$\text{Na}_2\text{C}_2\text{O}_4$

$\text{H}_2\text{C}_2\text{O}_4$

$M_1V_1=M_2V_2$

$M = n/V$

3. Calculate the molarity of dilute $\text{H}_2\text{C}_2\text{O}_4$

Drag the formula to be used here :

Type the final answer: mol L⁻¹

4. Calculate the molarity of stock solution $\text{H}_2\text{C}_2\text{O}_4$

Drag the formula to be used here :

Type the final answer: mol L⁻¹