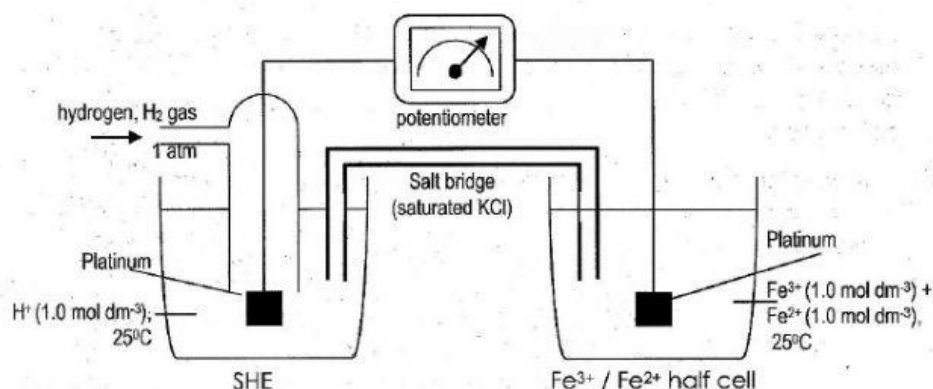


Fill in the blanks



$$E^{\circ} \text{ for } \text{H}^+/\text{H}_2 = 0\text{V}, E^{\circ} \text{ for } \text{Fe}^{3+}/\text{Fe}^{2+} = +0.77\text{V}$$

Oxidation occur at \_\_\_\_\_ and reduction occur at the \_\_\_\_\_. During the redox reaction,  $\text{H}_2$  is \_\_\_\_\_ to  $\text{H}^+$  and the oxidation number \_\_\_\_\_ from 0 to +1. While  $\text{Fe}^{3+}$  is \_\_\_\_\_ to  $\text{Fe}^{2+}$  and the oxidation number \_\_\_\_\_ from +3 to +2. The oxidising agent is \_\_\_\_\_ and the reducing agent is \_\_\_\_\_.

The observation is \_\_\_\_\_ solution turn into \_\_\_\_\_ solution

decrease	increase	$\text{H}_2$	oxidised	reduced
$\text{Fe}^{3+}/\text{Fe}^{2+}$ half cell	green	$\text{Fe}^{3+}$	SHE	yellow