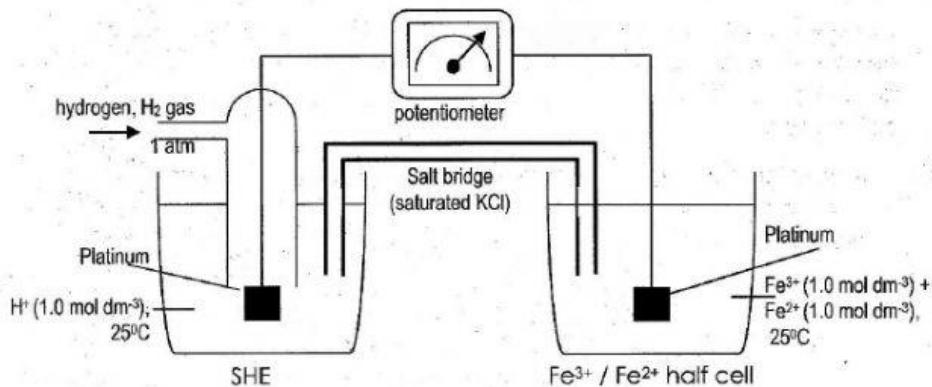


Fill in the blanks



$$E^\circ \text{ for } \text{H}^+/\text{H}_2 = 0\text{V}, E^\circ \text{ for } \text{Fe}^{3+}/\text{Fe}^{2+} = +0.77\text{V}$$

Oxidation occur at _____ and reduction occur at the _____. During the redox reaction, H₂ is _____ to H⁺ and the oxidation number _____ from 0 to +1. While Fe³⁺ is _____ to Fe²⁺ and the oxidation number from +3 to +2. The oxidising agent is _____ and the reducing agent is _____.

The observation is _____ solution turn into _____ solution

decrease	increase	H ₂	oxidised	reduced
Fe ³⁺ /Fe ²⁺ half cell	green	Fe ³⁺	SHE	yellow