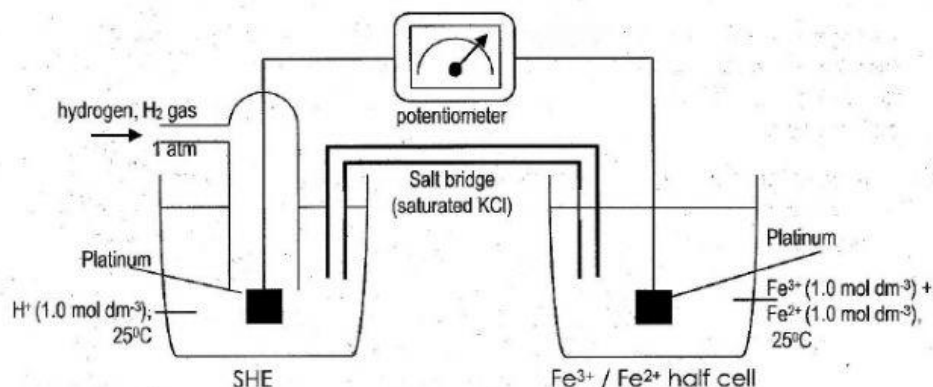


Fill in the blanks



$$E^\circ \text{ for } \text{H}^+/\text{H}_2 = 0\text{V}, E^\circ \text{ for } \text{Fe}^{3+}/\text{Fe}^{2+} = +0.77\text{V}$$

The half-equation at anode is $\rightarrow$

The half-equation at cathode is $\rightarrow$

The overall/ionic equation for the cell is

$\rightarrow$

The E_{cell} of the cell is

The cell diagram for the cell is

||

The electron flow from (SHE// $\text{Fe}^{3+}/\text{Fe}^{2+}$) to (SHE// $\text{Fe}^{3+}/\text{Fe}^{2+}$) half cell through external circuit.

$\text{Fe}^{2+}(\text{aq})$	$\text{H}_2(\text{g})$
$\text{Fe}^{3+}(\text{aq}), \text{Fe}^{2+}(\text{aq}) \text{Pt}(\text{s})$	$2\text{H}^+(\text{aq}) + 2\text{e}^-$
$\text{Fe}^{3+}(\text{aq}) + \text{e}^-$	$2\text{H}^+(\text{aq}) + 2\text{Fe}^{2+}(\text{aq})$
$\text{H}_2(\text{g}) + 2\text{Fe}^{3+}(\text{aq})$	$\text{Pt}(\text{s}) \text{H}_2(\text{g}) \text{H}^+(\text{aq})$