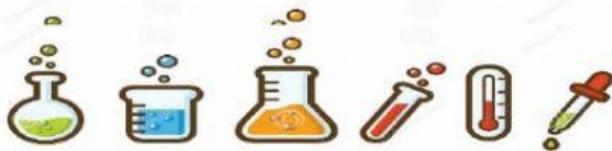


## CHEMBUDDY CHAPTER 3

### 3.1 CLASSIFICATION OF ELEMENTS



CHOOSE THE CORRECT ANSWER

NO	QUESTION	ANSWER
1	<p>What is group in periodic table?</p> <p>A. The horizontal row with label 1 to 7          B. The horizontal row with label 1 to 18          C. The vertical columns with label 1 to 18          D. The vertical columns with label 1 to 7 only</p>	A B C D
2	<p>What is period in periodic table?</p> <p>A. The horizontal row with label 1 to 7          B. The horizontal row with label 1 to 18          C. The vertical columns with label 1 to 18          D. The vertical columns with label 1 to 7 only</p>	A B C D
3	<p>Which block is NOT TRUE in periodic table?</p> <p>A. b block          B. d block          C. f block          D. p block</p>	A B C D
4	<p>Deduce period of element Y in periodic table. Given the electronic configuration of Y as <math>1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2</math></p> <p>A. period 1          B. period 2          C. period 3          D. period 4</p>	A B C D
5	<p>What is the charge on all the simple ions of elements of group 17?</p> <p>A. 1+          B. 2+          C. 1-          D. 2-</p>	A B C D
6	<p>Sodium (<math>Z=11</math>) and Silicon (<math>Z=14</math>) are examples of two elements which belong to the same</p> <p>A. group          B. period          C. block          D. compound</p>	A B C D

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	$^{37}_{17}P$ What is the position of this element in the periodic table?																										
7	<table> <thead> <tr> <th></th> <th>Block</th> <th>Group</th> <th>Period</th> <th></th> </tr> </thead> <tbody> <tr> <td>A.</td> <td>p</td> <td>17</td> <td>3</td> <td>A</td> </tr> <tr> <td>B.</td> <td>p</td> <td>15</td> <td>3</td> <td>B</td> </tr> <tr> <td>C.</td> <td>p</td> <td>7</td> <td>3</td> <td>C</td> </tr> <tr> <td>D.</td> <td>s</td> <td>5</td> <td>3</td> <td>D</td> </tr> </tbody> </table>		Block	Group	Period		A.	p	17	3	A	B.	p	15	3	B	C.	p	7	3	C	D.	s	5	3	D	
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8	$^{63}_{29}Q$ What is the position of this element in the periodic table? <table> <thead> <tr> <th></th> <th>Block</th> <th>Group</th> <th>Period</th> <th></th> </tr> </thead> <tbody> <tr> <td>A.</td> <td>s</td> <td>2</td> <td>4</td> <td>A</td> </tr> <tr> <td>B.</td> <td>s</td> <td>12</td> <td>4</td> <td>B</td> </tr> <tr> <td>C.</td> <td>d</td> <td>9</td> <td>3</td> <td>C</td> </tr> <tr> <td>D.</td> <td>d</td> <td>11</td> <td>4</td> <td>D</td> </tr> </tbody> </table>		Block	Group	Period		A.	s	2	4	A	B.	s	12	4	B	C.	d	9	3	C	D.	d	11	4	D	
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9	<p>An element R has a valence electronic configuration of <math>5s^2 4d^6</math>. What is the period and block of R in the Periodic Table?</p> <table> <thead> <tr> <th></th> <th>Block</th> <th>Period</th> <th></th> </tr> </thead> <tbody> <tr> <td>A.</td> <td>s</td> <td>5</td> <td>A</td> </tr> <tr> <td>B.</td> <td>s</td> <td>4</td> <td>B</td> </tr> <tr> <td>C.</td> <td>d</td> <td>5</td> <td>C</td> </tr> <tr> <td>D.</td> <td>d</td> <td>4</td> <td>D</td> </tr> </tbody> </table>		Block	Period		A.	s	5	A	B.	s	4	B	C.	d	5	C	D.	d	4	D						
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10	<p>An electronic configuration for element T is <math>1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4</math>. Determine the group of element T</p> <p>A. 4      B. 14      C. 6      D. 16</p>	A B C D																									
11	<p>An electronic configuration for element U is <math>1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^8</math>. Determine the group of element U.</p> <p>A. 8      B. 10      C. 14      D. 16</p>	A B C D																									

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12	The proton number for element V is 32. What is the valence orbital of element V? A. 3d 4p B. 4p C. 4s 4p D. 3d	A B C D
13	Elements in the Periodic Table are arranged in order increasing A. nucleon number B. proton number C. number of electron D. number of neutron	A B C D
14	Elements in the same group will have similar A. conductor properties B. thermal properties C. physical properties D. chemical properties	A B C D
15	Elements in Group 2 are known as A. alkaline earth metal B. noble gases C. alkali metal D. halogen	A B C D

