

CHOOSE THE CORRECT ANSWER

NO	QUESTION	ANSWER
1	The percent composition by mass of a compound is 76.0% C, 12.8% H, and 11.2% O. The molar mass of this compound is 284.5 g/mol. What is the molecular formula of the compound?	A. $C_9H_{18}O$ C. $C_{20}H_{12}O_2$ B. $C_{16}H_{28}O_4$ D. $C_{18}H_{36}O_2$
2	A 0.8715 g sample of sorbic acid, a compound first obtained from the berries of a certain ash tree, is burned completely in oxygen to give 2.053 g of carbon dioxide and 0.5601 g of water. The empirical formula of sorbic acid is	A. CH_2O C. CH_4O_3 B. C_3H_4O D. $C_3H_4O_2$
3	What volume of water in cm^3 should be added to 10.0 cm^3 of NaOH 6.0 M to produce a solution Of NaOH 0.3 M?	A. 10 C. 200 B. 190 D. 500
4	The label on the bottle of concentrated nitric acid solution is 37.4% HNO_3 by weight and density 1.18g/mL. Determine mole fraction of HNO_3 .	A. 0.374 C. 0.598 B. 0.171 D. 0.146
5	The density of 10.5 molal NaOH is 1.33 g/mL. Calculate percentage by mass of NaOH.	A. 12.67% C. 29.57% B. 75.18% D. 8.64%
6	A 3.2 molal solution contains 285 g of solvent. What is the molality of the solution after dilution with an additional 140 g of solvent?	A. 4.32 m C. 2.15 m B. 1.46 m D. 3.50 m
7	10 cm^3 of ammonia solution, NH_3 with density Of 0.93 g/ cm^3 contains 0.45 g of NH_3 solute. What is the percentage by mass of this solution?	A. 4.18 C. 2.22 B. 4.84 D. 4.20

8	A compound has an empirical formula of C_2H_7N . Choose the CORRECT statement.	<p>A. the weight percentage of carbon is 60.00 %</p> <p>B. the weight percentage of nitrogen is 31.11%</p> <p>C. the number of hydrogen atom is 5.21×10^{24}</p> <p>D. the number of nitrogen atoms is 3.0×10^{-24}</p>
9	The label on a commercial concentrated hydrochloric acid solution reads "37.4% HCl by weight and density 1.18 g mL^{-1} ". What is the molarity of HCl?	<p>A. 12.1 M</p> <p>B. 16.4 m</p> <p>C. 37.4 g/100mL</p> <p>D. 64.4 g/100mL</p>
10	A sample of gas contains 14 g of N_2 . This is equal to	<p>A. 1 mole of N_2</p> <p>B. 1 molar volume of N_2</p> <p>C. 2 moles of N_2</p> <p>D. 0.5 moles of N_2</p>