



- 1 Which of the following physical properties is true of ionic compounds?
- A Have high melting and boiling points
 - B Cannot conduct electricity in molten state
 - C Dissolve in organic solvents
 - D Exist as liquids at room temperature
- 2 A compound exists as a gas at room temperature. It dissolves in organic solvents. It cannot conduct electricity whether in solid or aqueous state. This compound could be
- A sodium chloride, NaCl .
 - B magnesium oxide, MgO .
 - C zinc bromide, ZnBr_2 .
 - D carbon dioxide, CO_2 .
- 3 An ionic compound X has very high melting point and boiling point because
- A the cations and anions are big.
 - B the electrostatic force between the cations and anions is very strong.
 - C the intermolecular force between molecules X is weak.
 - D the molecules in compound X cannot move freely.
- 4 An ionic compound Y cannot conduct electricity in solid state because
- A the cations and anions cannot move freely in solid state.
 - B the compound only contains neutral molecules in solid state.
 - C the ionic bond in solid state is weak.
 - D there are no ions in solid state.
- 5 The table below shows the proton numbers of elements P , Q , R and S .

Element	Proton number
P	3
Q	6
R	8
S	1

Which two elements will form a compound with high melting point and boiling point?

- A Q and R
- B S and Q
- C P and R
- D R and S