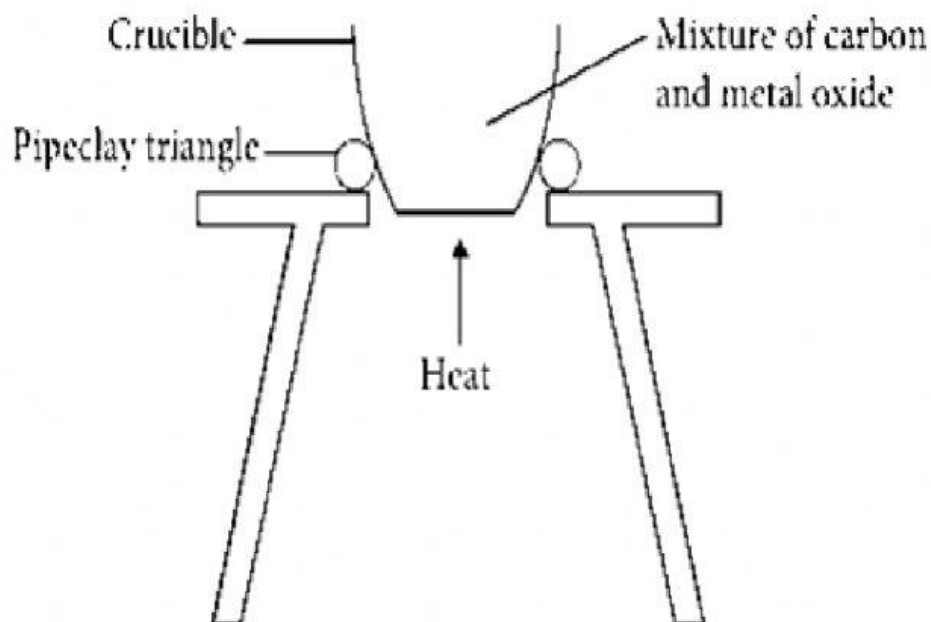


Position of Carbon in the Reactivity Series of Metals

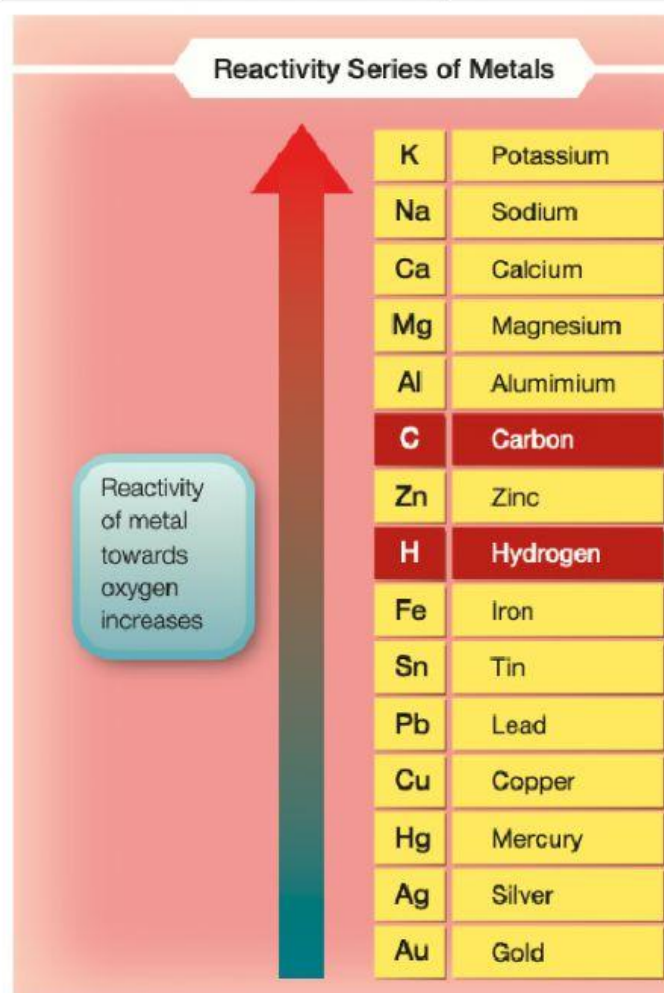
Activity 4.4: Determining the position of carbon in the reactivity series of metals.

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Observations

Mixture	Observation	Reactivity of carbon
Zinc oxide and carbon		
Aluminium oxide and carbon		
Lead(II) oxide and carbon		



1. Complete the word equation for each reaction of metal oxide with carbon, if any.

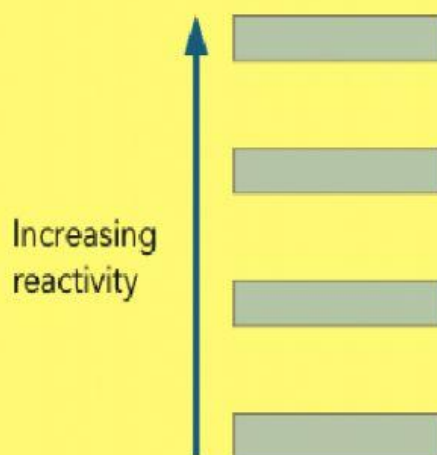
(a) Zinc oxide + carbon $\longrightarrow$

(b) Aluminium oxide + carbon $\longrightarrow$

(c) Lead(II) oxide + carbon $\longrightarrow$

2. Name the metal that is less reactive than carbon. Explain your answer.

3. Based on the results of this activity, complete the following sequence to show the arrangement of elements according to their increasing reactivity towards oxygen:



4. Give **one** application of the position of carbon in the reactivity series of metals for industrial use. Explain your answer.

5. Underline the correct answer for the following statements:

(a) If carbon can remove oxygen from a metal oxide, it means carbon is (more/less) reactive than the metal.

(b) If carbon cannot remove oxygen from a metal oxide, it means carbon is (more/less) reactive than the metal.