

# Chemistry – Atomic Structure

**Q. 1 Match the correct pairs : ( Use the pencil and draw lines )**

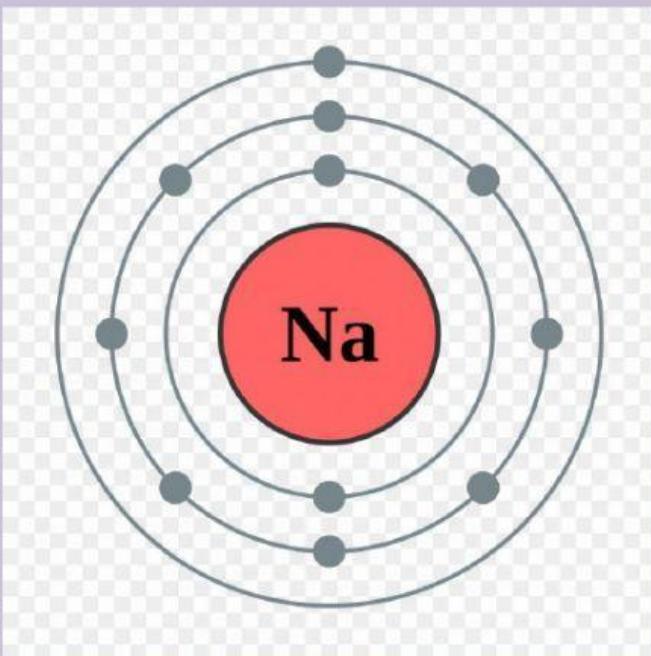
Element [protons]	Electronic Configuration
1. Magnesium[p=12]	a) 2, 8, 3
2. Sulphur[p=16]	b) 2, 8
3. Neon[p=10]	c) 2, 8, 2
4. Calcium[p=20]	d) 2, 8, 6
5. Aluminium[p=13]	e) 2, 8, 8, 2

**Q. 2 Choose the word options and drag them into the correct places:**

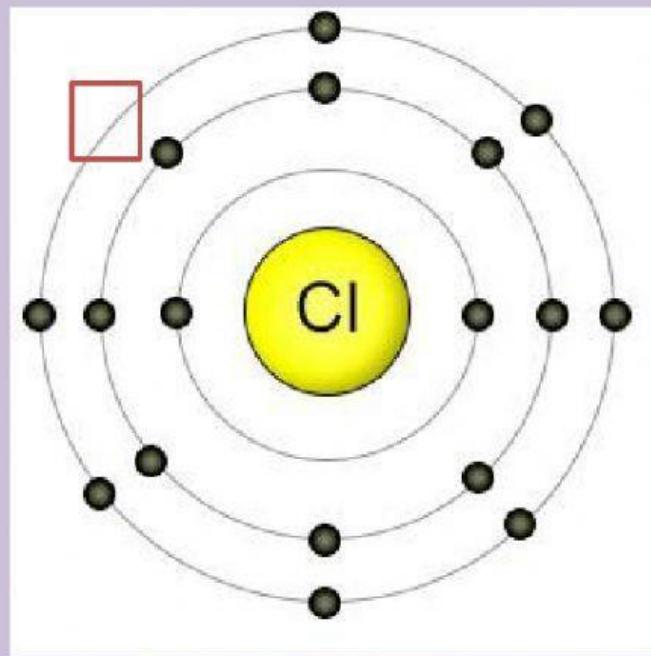
**2n<sup>2</sup>**      **Mass No.**      **Isotopes**      **Valency**      **Atomic No.**

- Atoms of the same element having the atomic number but different mass number.
- This is to tell us the total number of protons and neutrons present in a neutral atom.
- It is formula to determine the maximum number of electrons in each shell of an atom.
- This is to tell us the number of electrons or protons that are present in a neutral atom.
- This tells us the number of electrons donated or accepted by an atom to achieve a stable electronic configuration.

**Q. 3 Drag the electron/s from the first figure and drop it into the second figure to show the electron transfer and then answer the questions given below:**



**Fig 1 : SODIUM**



**Fig 2 : CHLORINE**

1. What kind of bond is achieved by this method? .....
2. Which of these two is a metal? .....
3. Which is the nearest inert gas to Chlorine? .....
4. Write the electronic configuration of Sodium. .....
5. What is the valency of Chlorine? .....

**Q. 4 Match the element symbol with its valency : (Draw lines )**

