

Chemistry – Atomic Structure

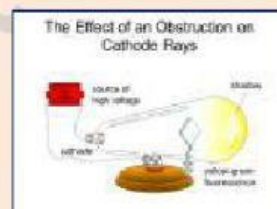
Q. 1 Match the correct pairs : (Use the pencil and draw lines)

Scientists

1. William Crooks
2. J.J. Thomson
3. E. Goldstein
4. Lord E. Rutherford
5. James Chadwick
6. Niels Bohr

Discoveries

- a) Electron orbit/shells
- b) Neutrons
- c) Atomic nucleus
- d) Protons
- e) Electrons
- f) The cathode rays



Q. 2 Choose the word options and drag them into the correct places:

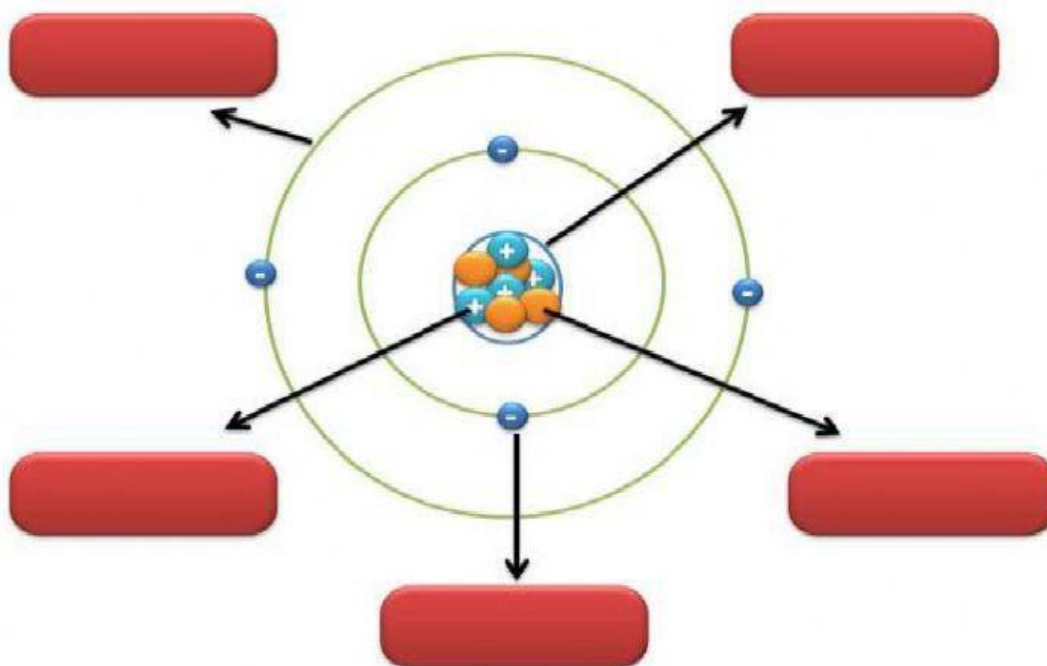
Electron

Orbit

Nucleus

Proton

Neutron



Q. 3 Complete the statements by typing the correct word from the brackets:

1. An element 'X' has six electrons in its outer or valence shell.
Its valency is [+2 / -2 / -1]
2. An element 'Y' has an electronic configuration 2,8,6.
The element 'Y' is a [metal / non-metal / noble gas]
3. A [proton / neutron] is a sub atomic particle with no charge and unit mass.
4. An element 'Z' with a zero valency is
[metal / non-metal / noble gas]
5. Magnesium atom with electronic configuration 2,8,2 achieves stable electronic configuration by losing two electrons, thereby achieving stable electronic configuration of the nearest noble gas [Neon / Argon]

Q. 4 Complete the blanks by choosing the correct option from the drop down menu:

1. Mass number of an atom is the number of protons and
2. The sub atomic particles with negligible mass
3. An atom having stable electronic configuration
4. a molecule by sharing of electrons (covalency).
5. A metallic atom having unstable electronic configuration