____ 11. The measure of the average kinetic energy of the particles of a substance is the _____.
a. kinetic energy
b. temperature
c. heat
d. thermal energy

____ 12. The change from a solid to a gas without becoming a liquid is called _____.
a. infiltration
b. evaporation
c. sublimation
d. dehydration

____ 13. The properties of solids, liquids, and gases are explained by the _____ theory.
a. conservation of mass
b. molecular
c. atomic
d. kinetic

____ 14. The energy required to change from the liquid state to the gaseous state is _____.
a. gaseous state energy
b. evaporation
c. boiling point
d. heat of vaporization

____ 15. The change from a solid to a liquid occurs at a temperature called the _____.
a. ice point
b. dew point
c. melting point
d. condensation point

____ 16. The temperature at which bubbles of vapor are formed within the liquid as it changes to a gas is the _____.
a. melting point
b. boiling point
c. dew point
d. sublimation point

Indentify

*For the following questions write in either **solid, liquid** or **gas.***

17. _____ Phase of matter with a definite shape and volume?
18. _____ Phase of matter that has no definite shape but a definite volume?
19. _____ Phase of matter with the most kinetic energy?
20. _____ Phase of matter with the least kinetic energy?
21. _____ Phase of matter with no definite shape and no definite volume?
22. _____ and _____ are the phases that exist at the melting point.
23. _____ and _____ are the phases that occur at the boiling point of a substance.

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