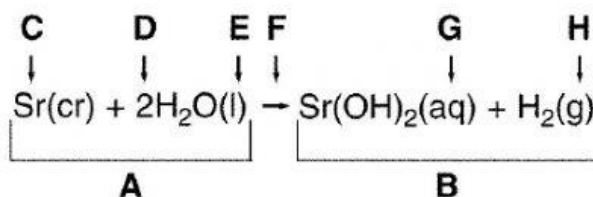


A) Label each of these parts of an equation. Use the word bank provided:

|                    |         |          |             |     |
|--------------------|---------|----------|-------------|-----|
| Yield or Produce   | Product | Reactant | Coefficient | Gas |
| Dissolved in Water | Liquid  | Solid    | Element     |     |



- 1) In the chemical equation shown, the substances in part A are called \_\_\_\_\_.
- 2) Part F is read as \_\_\_\_\_.
- 3) The substances in part B are called \_\_\_\_\_.
- 4) Part G means that the material,  $\text{Sr(OH)}_2$ , is \_\_\_\_\_.
- 5) Part D is called a(n) \_\_\_\_\_.
- 6) Part E means that the material,  $\text{H}_2\text{O}$ , is a(n) \_\_\_\_\_.
- 7) Part H means that the material,  $\text{H}_2$ , is a(n) \_\_\_\_\_.
- 8) Part C is a(n) \_\_\_\_\_.

B) 1) When heated, mercury (II) oxide produces oxygen plus mercury. What would be the combined mass of oxygen and mercury if 20 g of mercury (II) oxide were heated?

2) How much oxygen would be produced by heating 20 g of mercury (II) oxide if 18.6 g of mercury is produced?

C) Which type of chemical reaction is each of these? Use these terms:

|           |               |                     |                     |            |
|-----------|---------------|---------------------|---------------------|------------|
| Synthesis | Decomposition | Single Displacement | Double Displacement | Combustion |
|-----------|---------------|---------------------|---------------------|------------|

