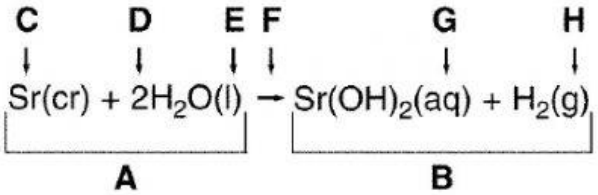
ETIVEMORKSHEETS

A) Label each of these parts of an equation. Use the word bank provided:

Yield or Produce	Product	Reactant	Coefficient	Gas
Dissolved in Water	Liquid	Solid	Element	



1) In the chemical equation shown, the substances in part A are called _____.

2) Part F is read as _____.

3) The substances in part B are called _____.

4) Part G means that the material, $Sr(OH)_2$, is _____.

5) Part D is called a(n) _____.

6) Part E means that the material, H_2O, is a(n) _____.

7) Part H means that the material, H_2, is a(n) _____.

8) Part C is a(n) _____.

B) 1) When heated, mercury (II) oxide produces oxygen plus mercury. What would be the combined mass of oxygen and mercury if 20 g of mercury (II) oxide were heated?

2) How much oxygen would be produced by heating 20 g of mercury (II) oxide if 18.6 g of mercury is produced?

C) Which type of chemical reaction is each of these? Use these terms:

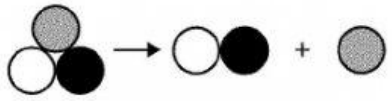
Synthesis Decomposition Single Displacement Double Displacement Combustion

a. $Mg + O_2 \rightarrow MgO$ _____

b. $Pb(NO_3)_2 + KI \rightarrow PbI_2 + KNO_3$ _____

c. $CH_4 + O_2 \rightarrow CO_2 + 2H_2O$ _____

d.  _____

e.  _____

f. $KClO_3 \rightarrow KCl + O_2$ _____