**I. Multiple Choice Questions. (One-mark questions.)**

1. An example for chemical change among the following is (A)
   A. Dissolving salt in water B. Tearing of paper
   C. Rusting of iron D. Boiling of water.
2. This is not a property of a chemical reaction among the following. (A)
   A. Change in colour B. Release of heat
   C. Change in shape D. Release of gas.
3. $Fe_2O_3 + 2Al \rightarrow Al_2O_3 + 2Fe$ (A)
   The above chemical reaction is an example for
   A. Combination reaction B. Decomposition reaction
   C. Displacement reaction D. Double displacement reaction.
4. The group of reactants among the following which react with each other to exchange ions and form a precipitate is (D)
   A. $BaCl_2$ and $Na_2SO_4$ B. $Al_2O_3$ and HCl
   C. NaOH and $H_2SO_4$ D. $Na_2O$ and $CO_2$
5. The evolution of Oxygen and Hydrogen gases in the electrolysis of water is (E)
   A. Double displacement reaction B. Exothermic reaction
   C. Displacement reaction D. Decomposition reaction
6. The chemical reaction that takes place when reactants exchange their ions mutually is (A)
   A. Double displacement reaction B. Dissociation reaction
   C. Dissociation reaction D. Combination reaction
7. $CuO + H_2 \rightarrow Cu + H_2O$ (D)
   In this reaction, the reactants that undergo oxidation and reduction respectively are
   A. $CuO + H_2$ B. $H_2 + CuO$ C. $Cu + H_2O$ D. $H_2O + Cu$
8. Silver nitrate turns into grey colour in the presence of sunlight because (D)
   A. Silver chloride decomposes to form silver.
   B. Silver chloride decomposes to form chlorine.
   C. Silver chloride undergoes oxidation.
   D. Silver chloride undergoes reduction.
9. The ratio of hydrogen and oxygen gases released in the electrolysis of water is (D)
   A. Hydrogen : Oxygen:: 1 : 2 B. Oxygen : Hydrogen:: 2 : 3
   C. Hydrogen : Oxygen :: 2 : 1 D. Oxygen : Hydrogen:: 3 : 2
10. The conversion of vegetable waste into compost is an example for (A)
    A. Reduction reaction B. Exothermic reaction
    C. Endothermic reaction D. Redox reaction

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