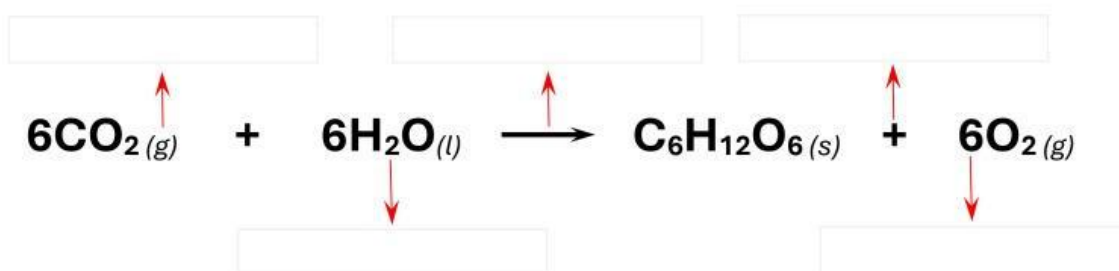


Review CHAPTER 10

1. Label the components of the Chemical Reaction below by selecting the correct options from the drop-down boxes :



2. Draw a line between the options on the left and the matching options on the right :

Products	Starting Materials
Precipitate	Number of Atoms
Subscript	Number of Molecules
Reactants	Newly formed Substances
Coefficient	Solid that forms in a liquid

3. Identify the type of Chemical Reaction in each of the following scenarios (select only 1 option in each case unless stated otherwise) :

3.1) $\text{Na} + \text{Cl} \rightarrow \text{Table salt}$

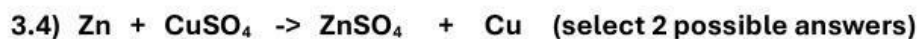
- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Double Replacement
- E) Combustion

3.2) $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$

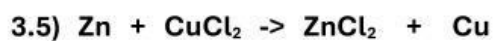
- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Double Replacement
- E) Combustion



- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Double Replacement
- E) Combustion



- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Combustion
- E) Redox



- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Double Replacement
- E) Combustion



- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Double Replacement
- E) Combustion



- A) Synthesis
- B) Decomposition
- C) Single Replacement
- D) Double Replacement
- E) Combustion

4. Drag and Drop the words from the Word Bank onto the correct places so that the statements below make sense :

gained reactant decrease product increase lost

In a Synthesis Reaction, 2 or more reactants are combined into a single _____ .

Oxygen is always a _____ in a Combustion Reaction.

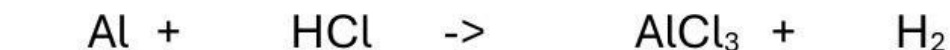
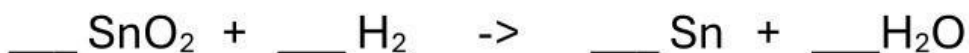
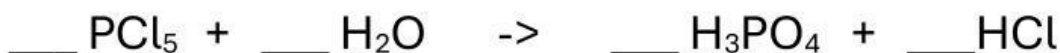
In a Redox Reaction, the element that undergoes Reduction will have a(n) _____ in

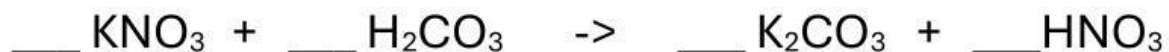
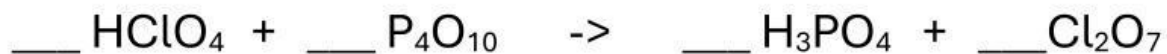
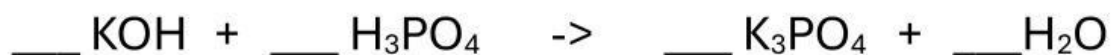
Oxidation Number. This means that this element has _____ one or more Electrons.

The element that undergoes Oxidation, will have a(n) _____ in Oxidation Number.

This means that this element has _____ one or more Electrons.

5. Balance the following Chemical Equations (The grading tool unfortunately cannot grade empty spaces, so please do NOT leave any spaces empty ... for the sake of the grading tool, please fill in the "1" in those spaces) :





6. Look at the Chemical Formulas provided and answer the questions that follow :

6.1) How many Carbon atoms are there in $\text{C}_4\text{H}_2\text{S}(\text{OH})_2$: _____

6.2) How many Carbon atoms are there in $3\text{C}_4\text{H}_2\text{S}(\text{OH})_2$: _____

6.3) How many Hydrogen atoms are there in $\text{C}_4\text{H}_2\text{S}(\text{OH})_2$: _____

6.4) How many Hydrogen atoms are there in $3\text{C}_4\text{H}_2\text{S}(\text{OH})_2$: _____

6.5) How many Oxygen atoms are there in $2\text{C}_4\text{S}(\text{OH})_2$: _____

6.6) How many Hydroxide atoms are there in $2\text{C}_4\text{S}(\text{OH})_2$: _____