Cl = 35.45 a.m.u
K= 39 a.m.u
O = 16 a.m.u

Q1 : $2\ KClO_3 \rightarrow 2\ KCl + 3\ O_2$

How many moles of O_2 will be formed from 1.65 moles of $KClO_3$?

$$2.73 \text{ moles of } KClO_3 \text{ X } \frac{\text{mol } O_2}{\text{mol } KClO_3} = \text{mol } O_2$$

Q2 : How many **gram** of KCl will be formed from 2.73 moles of $KClO_3$?

2.73 moles of $KClO_3$ X _____ X _____ = **gram** of KCl

LIVEWORKSHEETS