

|                             | Symbols & Oxidation #                                                                    | Formula                                                             | Name of Compound |
|-----------------------------|------------------------------------------------------------------------------------------|---------------------------------------------------------------------|------------------|
| 16. Sodium and phosphate    | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 17. Ammonium and hydroxide  | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 18. Calcium and nitrate     | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 19. Beryllium and hydroxide | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 20. Potassium and acetate   | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 21. Sodium and carbonate    | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 22. Sodium and bicarbonate  | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |
| 23. Aluminum and sulfate    | <div><div></div><div><div></div><div></div></div><div><div></div><div></div></div></div> | <div><div></div><div></div></div> <div><div></div><div></div></div> | <div></div>      |

25. An ion is an atom or group of atoms that has become electrically

26. When an atom loses an electron its charge is **(positive or negative)**

27. An ionic bond is the attraction between **(opposites, positive, neutral, or negative)** ions.

28. Ionic compounds are electrically **(charged, positive, neutral, or negative)**.

29. The sum of the charges for an ionic compound is  

The two answers must be in the right order.

30. An ionic compound is the result of the bonding of a **(non-metal, metalloid, metal, noble gas)** with a **(non-metal, metalloid, metal)**.