

PART A

REACTION OF STRONG ACID AND STRONG BASE



1. Drag the acid and base to form the salt produced.
2. On a piece of paper, write equation of dissociation for the salt.

Identify the correct answer:

K^+ : strong conjugate acid Br^- : strong conjugate base	Both ions can hydrolyse	Code : 3
K^+ : weak conjugate acid Br^- : weak conjugate base	Both ions do not hydrolyse	Code : 1
K^+ : weak conjugate base Br^- : weak conjugate acid	K^+ ions can hydrolyse Br^- ions do not hydrolyse	Code : 2
K^+ : strong conjugate acid Br^- : strong conjugate base	Both ions do not hydrolyse	Code : 5

On the same paper, write equations of the ions with water (if any).

Classify the salt formed:

Neutral / acidic / basic

PART B

REACTION OF STRONG ACID AND WEAK BASE



1. Drag the acid and base to form the salt produced.
2. On a piece of paper, write equation of dissociation for the salt.

Identify the correct answer:

NH_4^+ : weak conjugate acid Br^- : strong conjugate base	Both ions can hydrolyse	Code : 1
NH_4^+ : weak conjugate acid Br^- : weak conjugate base	Both ions do not hydrolyse	Code : 0
NH_4^+ : weak conjugate base Br^- : weak conjugate acid	NH_4^+ ions do not hydrolyse Br^- ions can hydrolyse	Code : 3
NH_4^+ : strong conjugate acid Br^- : weak conjugate base	NH_4^+ ions can hydrolyse Br^- ions do not hydrolyse	Code : 5

On the same paper, write equations of the ions with water (if any).

Classify the salt formed:

Neutral / acidic / basic

PART C

REACTION OF STRONG BASE AND WEAK ACID



1. Drag the acid and base to form the salt produced.
2. On a piece of paper, write equation of dissociation for the salt.

Identify the correct answer:

Na^+ : weak conjugate acid CH_3COO^- : strong conjugate base	Na^+ ions do not hydrolyse CH_3COO^- ions can hydrolyse	Code : 3
Na^+ : weak conjugate acid CH_3COO^- : weak conjugate base	Both ions do not hydrolyse	Code : 1
Na^+ : weak conjugate base CH_3COO^- : weak conjugate acid	Na^+ ions do not hydrolyse CH_3COO^- ions can hydrolyse	Code : 9
Na^+ : strong conjugate acid CH_3COO^- : weak conjugate base	Na^+ ions can hydrolyse CH_3COO^- ions do not hydrolyse	Code : 7

On the same paper, write equations of the ions with water (if any).

Classify the salt formed:

Neutral / acidic / basic