PART A

REACTION OF STRONG ACID AND STRONG BASE



1. Drag the acid and base to form the salt produced.
2. On a piece of paper, write equation of dissociation for the salt.

Identify the correct answer:

K^+: strong conjugate acid Br^-: strong conjugate base	Both ions can hydrolyse	Code : 3
K^+: weak conjugate acid Br^-: weak conjugate base	Both ions do not hydrolyse	Code : 1
K^+: weak conjugate base Br^-: weak conjugate acid	K^+ ions can hydrolyse Br^- ions do not hydrolyse	Code : 2
K^+: strong conjugate acid Br^-: strong conjugate base	Both ions do not hydrolyse	Code : 5

On the same paper, write equations of the ions with water (if any).

Classify the salt formed:

Neutral / acidic / basic

LIVEWORKSHEETS

PART B

REACTION OF STRONG ACID AND WEAK BASE



1. Drag the acid and base to form the salt produced.
2. On a piece of paper, write equation of dissociation for the salt.

Identify the correct answer:

NH_4^+: weak conjugate acid Br^-: strong conjugate base	Both ions can hydrolyse	Code : 1
NH_4^+: weak conjugate acid Br^-: weak conjugate base	Both ions do not hydrolyse	Code : 0
NH_4^+: weak conjugate base Br^-: weak conjugate acid	NH_4^+ ions do not hydrolyse Br^- ions can hydrolyse	Code : 3
NH_4^+: strong conjugate acid Br^-: weak conjugate base	NH_4^+ ions can hydrolyse Br^- ions do not hydrolyse	Code : 5

On the same paper, write equations of the ions with water (if any).

Classify the salt formed:

Neutral / acidic / basic

LIVEWORKSHEETS

PART C

REACTION OF STRONG BASE AND WEAK ACID

CH_3COO H + Na OH → _____ + H_2O

1. Drag the acid and base to form the salt produced.
2. On a piece of paper, write equation of dissociation for the salt.

_____

Identify the correct answer:

Na^+: weak conjugate acid CH_3COO^-: strong conjugate base	Na^+ ions do not hydrolyse CH_3COO^- ions can hydrolyse	Code : 3
Na^+: weak conjugate acid CH_3COO^-: weak conjugate base	Both ions do not hydrolyse	Code : 1
Na^+: weak conjugate base CH_3COO^-: weak conjugate acid	Na^+ ions do not hydrolyse CH_3COO^- ions can hydrolyse	Code : 9
Na^+: strong conjugate acid CH_3COO^-: weak conjugate base	Na^+ ions can hydrolyse CH_3COO^- ions do not hydrolyse	Code : 7

On the same paper, write equations of the ions with water (if any).

Classify the salt formed:

Neutral / acidic / basic

LIVEWORKSHEETS