

## Bellringers Week 8: 10/13 - 10/16

### Day 1: Monday 10/13

- How many electrons can fit into each of these sublevels:

s:      p:      d:      f:

- Use your periodic table blocks and tell me how many and which orbitals are on each of the following energy levels: for example (1, s)

Ring/Energy Level/Orbital 1:

Ring/Energy Level/Orbital 2:

Ring/Energy Level/Orbital 3:

Ring/Energy Level/Orbital 4:

Ring/Energy Level/Orbital 5:

Ring/Energy Level/Orbital 6:

Ring/Energy Level/Orbital 7:

### Day 2: Tuesday 10/14

- Use shorthand electron configuration to write out the electron configurations for these elements:

<sup>38</sup>Sr:

<sup>33</sup>As:

<sup>63</sup>Eu:

<sup>74</sup>W:

### Day 3: Wednesday 10/15

- PSAT Day

### Day 4: Thursday 10/16

Match each rule to its definition:

|                           |                                                                            |
|---------------------------|----------------------------------------------------------------------------|
| Aufbau principle          | orbitals can hold a maximum of two electrons with opposite spins           |
| Hund's rule               | Fill orbitals from lowest to highest energy                                |
| Pauli exclusion principle | electrons first occupy each orbital singly before any orbital is paired up |

**Day 5: Friday 10/17**

- NO SCHOOL