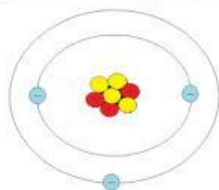


## Atomic Structure Review

### **I. Fill in the chart:**

| Element    | Atomic # | Mass # | Protons | Neutrons | Electrons |
|------------|----------|--------|---------|----------|-----------|
| Technetium |          |        |         |          |           |
| Tellurium  |          |        |         |          |           |
| Oxygen     |          |        |         |          |           |
| Xenon      |          |        |         |          |           |
| Iodine     |          |        |         |          |           |
| Carbon     |          |        |         |          |           |
| Sulfur     |          |        |         |          |           |
| Nitrogen   |          |        |         |          |           |
| Chromium   |          |        |         |          |           |
| Selenium   |          |        |         |          |           |

### **II. Which elements are represented by these maps:**



**III. Answer these Questions:** Use your notes to answer the following questions.

1. Who named the atom?

2. Which model is called the plum pudding model?

3. What are electrons equal to?

4. What is the maximum number of energy levels an atom can have?

5. What is the overall charge of an atom?

6. What did Thomson discover?

7. Which model is called the planetary model?

8. What percentage of the alpha particles in Rutherford's experiment went straight through the gold foil?

9. Who discovered that the atom is mostly empty space?

10. Who named the energy levels?

11. How many electrons can energy level 2 hold?

12. What year was the electron cloud model designed?

13. What is the smallest part of an element?

14. What did Schrodinger discover about the atom?

15. What is the name of Dalton's theory?

16. What is the name of the model we use today to describe the atom?

17. How do you find the number of neutrons in an atom?

18. What tells you the number of protons in an atom?

19. What is the sum of protons and neutrons called?

20. How many electrons can be held in energy level 4?

21. How many electrons can be held in energy level 3?

22. Write the 4 parts of the atomic theory.

- 1.
- 2.
- 3.
- 4.

23. Does the atomic number of an element ever change?

24. How do you find the number of electrons in an atom?

25. What are the building blocks of all matter?

26. Which particle in the atom has a neutral charge?

27. What is a proton?

28. What is an electron?

29. Why are atoms on the periodic table neutral?

30. What is the atomic number?

31. Define energy level.

32. Define atom.

33. What does the map of an atom show?

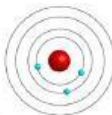
IV. Whose Diagram?: Write the name of the scientist or model in the box.



34.



35.



36.



37.



38.

