Questions Must Be Completed On Separate Paper

Analysis and Conclusions:

1. Define oxidation: _____
2. Define reduction: _____
3. What is the name of the gas produced by the reactions in Step 1?_____
4. a. Write the equation for the reaction between hydrochloric acid and magnesium

   b. Label the oxidation numbers of each element in the equation above.

   c. Which substance is oxidized? _____

   d. Which substance is reduced? _____
5. a. Write the equation for the reaction between hydrochloric acid and zinc.

   b. Label the oxidation numbers of each element in the equation above.

   c. Which substance is oxidized? _____

   d. Which substance is reduced? _____
6. Based on the reactions in Steps 2 and 3, which is more easily reduced, Cu or Zn? _____
7. a. Write the equation for the reaction between copper nitrate and zinc.

   b. Label the oxidation numbers of each element in the equation above.

   c. Which substance is oxidized? _____

   d. Which substance is reduced? _____