



THIRD QUARTERLY ASSESSMENT
Science 7
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Name: _____ Score: _____ /70 _____

Grade and Section: _____ Date: _____

I. A. Identify what is being described.

PHYSICAL CHANGE	ELEMENT	LIQUID	
CHEMICAL CHANGE	MATTER	SOLID	PURE SUBSTANCE
COMPOUND	GAS	MOLECULES	

_____ 1. A substance made up of various types of particles that occupies physical space and has inertia.

_____ 2. Phase of matter without a definite shape and definite volume. The particles are far from each other.

_____ 3. Phase of matter with definite shape and definite volume, and composed of particles which are bonded together.

_____ 4. Composed of particles which are held together by forces that are not that strong as compared to those in solids.

_____ 5. A process which does not produce a new product and objects retained their identities.

_____ 6. A process that produces a new substance, identities and the composition of the original substance is changed.

_____ 7. Homogeneous material consists of only one kind of material with unique set of physical and chemical properties.

_____ 8. Made up of two or more elements present in specific ratio wherein the properties are totally different from the elements.

_____ 9. It is the smallest unit of an element containing the properties of that particular element.

_____ 10. A combination of different elements. It contains two oxygen atoms bonded together.

B. Name the elements based on their chemical symbol below.

11. **Be** → _____

12. **Mn** → _____

13. **Sg** → _____

14. **Rb** → _____

15. **Zr** → _____

16. **Pd** → _____

17. **K** → _____

18. **Ca** → _____

19. **Mg** → _____

20. **He** → _____

II. Solve for the Law of Conservation of Mass. Fill in the blanks.

$\text{NaCl} \rightarrow \text{Na} + \text{(21.)}$ _____

64g. **(22.)** _____ **(23.)** _____

$\text{SO}_2 \rightarrow \text{(24.)}$ _____ + SO

32g. 16g. 16g

$\text{H} + \text{O} \rightarrow \text{(25.)}$ _____

17g 17g. **(26.)** _____

$\text{H}_2\text{O} \rightarrow \text{(27.)}$ _____ + **(28.)** _____

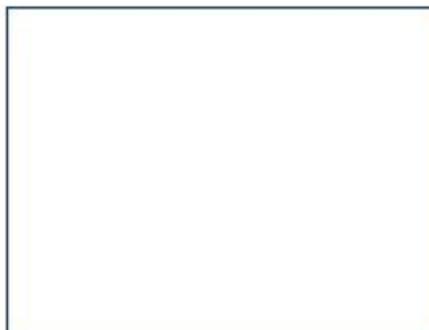
14g. 7g. 7g

$\text{Fe} + \text{C} \rightarrow \text{(29.)}$ _____

14g. 14g. **(30.)** _____

III. Draw and explain the idea of each atomic model.

31-33. Nuclear Model



34-36. Planetary Model



37-39. Plum Pudding Model



40-42. Quantum Model



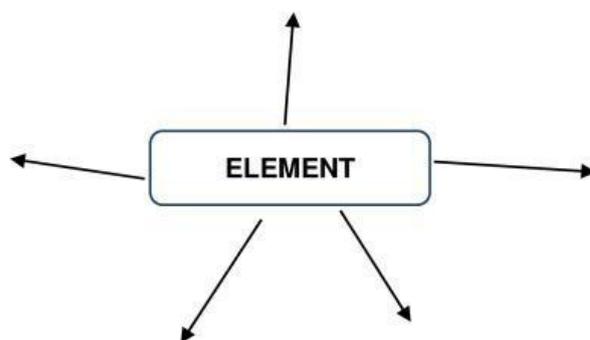
43-45. Sphere Model



IV. Solve for the Law of Definite Composition. Complete the table below and show your solution.

Case	Sulfur	Oxygen	Percentage
1	2.01	7.65	46-47.)
2	14.25	21	48-49.)
3	5.07	14.25	50-51.)
4	7.65	10.01	52-53.)
5	7	21	54-55.)

V. Complete the concept map below. Write 5 ideas about an Element. (56-60.)



VI. Explain the following in 2-3 sentences.

61-62. What is the importance of studying Matter?

63-64. How do different types of matter affect the environment, technology, and health? Provide an example.

65-66. Why do we have to study the elements in the Periodic Table?

67-68. Describe the three main states of matter: Solid, Liquid, and Gas. How do they differ from each other?

69-70. Choose 2 Elements from the periodic table and discuss their importance in everyday life.

***** END OF EXAM *****