

Name _____
Chemistry _____

Date _____
Aversano/Herz _____

Student Study Guide for Chemical Reactions Test

Directions: Use your notes and chemistry reference table to help you complete this study guide.

Key Topics to Study

Types of Reactions

1. Synthesis Reaction

- Two or more substances combine to form a single product.
- General Formula: $A + B \rightarrow AB$
- Example: _____

2. Decomposition Reaction

- A single compound breaks down into two or more products.
- General Formula: $AB \rightarrow A + B$
- Example: _____

3. Single Replacement Reaction

- One element replaces another in a compound.
- General Formula: $A + BC \rightarrow AC + B$
- Example: _____

4. Double Replacement Reaction

- The ions of two compounds exchange places in an aqueous solution.
- General Formula: $AB + CD \rightarrow AD + CB$
- Example: _____

5. Combustion Reaction

- A hydrocarbon reacts with oxygen to produce carbon dioxide and water.
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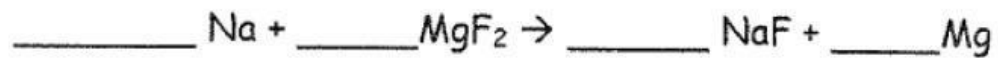
Practice Questions:

1. Classify the reaction: $\text{NaCl} + \text{AgNO}_3 \longrightarrow \text{AgCl} + \text{NaNO}_3$

2. Identify the reaction type: $\text{CuCO}_3 \longrightarrow \text{CuO} + \text{CO}_2$

3. Identify the reaction type: $\text{H}_2 + \text{N}_2 \longrightarrow \text{NH}_3$

4. Balance and classify the reaction:



Formula Writing and Naming

- Ionic Compounds:
 - Combine cation (positive ion) and anion (negative ion).
 - Use Roman numerals for transition metals to indicate charge.
 - Example: = Iron (III) chloride.

Practice Questions:

1. Name: Sodium Sulfate

2. Write the formula for nitrogen oxide.

3. Write the formula Copper (II) Nitrate.

4. Write the formula for sulfur fluoride

5. Name: KClO_3

Empirical Formula

- The simplest whole-number ratio of elements in a compound.
- Steps to Determine:
 1. Convert mass (or %) to moles.
 2. Divide by the smallest number of moles.
 3. Multiply to get whole numbers if necessary.

Practice Questions:

1. Determine the empirical formula for N_4O_6

2. Determine the empirical formula for C_9H_{18}

Mass Percent

- The percentage by mass of an element in a compound.

- Formula: _____

Example:

- Find the mass % of hydrogen in H_2O

Practice Questions:

1. Calculate the mass percent of oxygen in _____CO₂_____
2. Find the mass percent of chlorine in HClO₃
3. What is the mass percent of carbon in NaHCO₃ ?
4. Determine the mass percent of sodium in Na₂SO₄.

Converting Moles to Grams

- Formula: _____

Example:

- Convert 2 moles of NaCl to grams:

Practice Questions:

1. How many grams are in 0.75 moles of KI?

2. How many grams are in 4.5 moles of CO?