

Weekly Warm Ups

Honors Chemistry

Monday

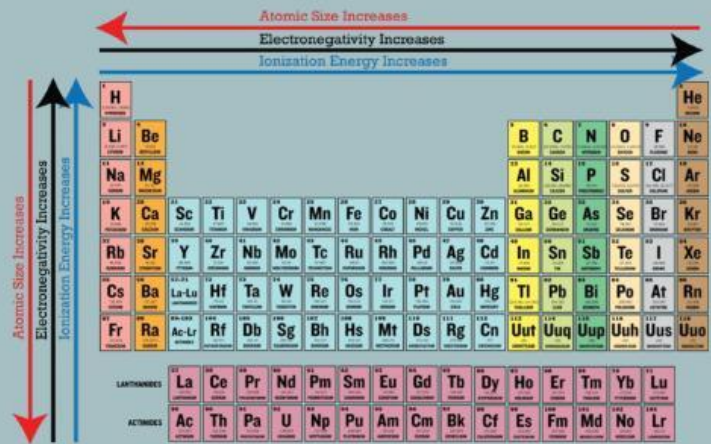
Tuesday

Wednesday

The octet rule explains the stability of most covalently bonded molecules in terms of:

- A. ? noble-gas configurations
- B. ? ionization energies
- C. ? electronegativity differences
- D. ? electron affinities

Which element is the
MOST
electronegative?



Thursday

In drawing Lewis structures, a single line (single bond) between two elements represents:

- A. ☐ a shared electron
- B. ☐ an octet of electrons
- C. ☐ an unshared pair of electrons
- D. ☐ a shared pair of electrons

Which of the following is an acceptable Lewis Structure for the diatomic nitrogen molecule?

- A. ☐ $\ddot{\text{N}}=\ddot{\text{N}}$
- B. ☐ $:\text{N}\equiv\text{N}:$
- C. ☐ $\cdot\text{N}\equiv\text{N}\cdot$
- D. ☐ $\ddot{\text{N}}\equiv\ddot{\text{N}}$

Which of the diatomic elements has a double bond between its atoms?

- A. ☐ fluorine
- B. ☐ nitrogen
- C. ☐ hydrogen
- D. ☐ oxygen

Friday

The seven elements that occur as diatomic elements are:

- A. ☐ $\text{Fe}_2, \text{Rn}_2, \text{O}_2, \text{He}_2, \text{Ne}_2, \text{C}_2, \text{Br}_2$
- B. ☐ $\text{H}_2, \text{N}_2, \text{O}_2, \text{He}_2, \text{Ne}_2, \text{Cl}_2, \text{Br}_2$
- C. ☐ $\text{H}_2, \text{N}_2, \text{O}_2, \text{He}_2, \text{Ne}_2, \text{C}_2, \text{Na}_2$
- D. ☐ $\text{H}_2, \text{N}_2, \text{O}_2, \text{F}_2, \text{Cl}_2, \text{Br}_2, \text{I}_2$

In the correct Lewis structure for water, how many unshared pairs of electrons will oxygen have?

- A. ☐ 1
- B. ☐ 2
- C. ☐ 3
- D. ☐ 4