

## DO NOW-MINI LAB #1

You have a balloon and a test tube on your desk. The test tube has vinegar [ $\text{HC}_2\text{H}_3\text{O}_2$ ] in it, the balloon has baking soda in it [ $\text{NaHCO}_3$ ]. Pour the Baking Soda into the test tube and observe.

The reaction equation is:

1. What are the reactants? (Select the FORMULAS)  
  $\text{NaHCO}_3$    $\text{HC}_2\text{H}_3\text{O}_2$    $\text{H}_2\text{O}$    $\text{CO}_2$    $\text{NaC}_2\text{H}_3\text{O}_2$
2. What are the products? (Select the FORMULAS)  
  $\text{NaHCO}_3$    $\text{HC}_2\text{H}_3\text{O}_2$    $\text{H}_2\text{O}$    $\text{CO}_2$    $\text{NaC}_2\text{H}_3\text{O}_2$
3. What gas is produced? (Write the formula)

4. What liquid produced? (Write the formula)

5. Is the reaction balanced? YES NO

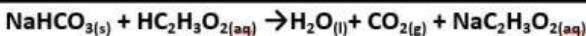
6. If the equation is not balanced, what's wrong and how could you fix it? Select where you would add coefficients

$\text{NaHCO}_3$    $\text{HC}_2\text{H}_3\text{O}_2$    $\text{H}_2\text{O}$    $\text{CO}_2$    $\text{NaC}_2\text{H}_3\text{O}_2$

7. Is the reaction (exothermic) letting heat out, or (endothermic) absorbing heat?

8. How can you tell this was a chemical change and not a physical change?  
You can tell because there was a \_\_\_\_\_ produced.

9. The substance that settled to the bottom of the test tube is called a \_\_\_\_\_



Na

H

C

O