

1+	1	
1	2+	Oxidation Number
H	-2	Valence Electrons
Hydrogen 1.008	2	Family
3	4	
Li	Be	
Lithium 6.941	Beryllium 9.0122	
11	12	
Na	Mg	
Sodium 22.990	Magnesium 24.305	
19	20	
K	Ca	
Potassium 39.098	Calcium 40.078	

Oxidation Number	3+	4+/-	3-	2-	1-	0
Valence Electrons	3	4	5	6	7	8
Family	3	4	5	6	7	He
	Boron 10.81	Carbon 12.011	Nitrogen 14.007	Oxygen 15.998	Fluorine 18.998	Neon 20.180
	Aluminum 26.98	Silicon 28.086	Phosphorus 30.974	Sulfur 32.06	Chlorine 35.453	Argon 39.948
	Gallium 69.72	Germanium 72.81	Arsenic 74.922	Selenium 78.90	Bromine 79.904	Krypton 83.80

Some Polyatomic Ions and their Oxidation Numbers				
1+	1-	2-	3-	
ammonium (NH ₄)	acetate (C ₂ H ₃ O ₂)	carbonate (CO ₃)	sulfate (SO ₄)	phosphate (PO ₄)
	chlorate (ClO ₃)			
	hydroxide (OH)			
	nitrate (NO ₃)			
	bicarbonate(HNO ₃)			

8. The oxidation number of an ion is the number of electrons an atom will lose to become a negative positive ion or gain to become a negative positive ion and have a full outer energy level of electrons.

Combine the elements or polyatomic compounds and write the formula and the name

	Symbols & Oxidation #	Formula	Name of Compound
9. Sodium and phosphate			
10. Magnesium and oxygen			
11. Potassium and phosphorus			
12. Ammonium and hydroxide			
13. Sodium and chlorine			
14. Beryllium and bromine			