

1+	1	1	2+	Oxidation Number	3+	4+/-	3-	2-	1-	0
1	1	1	2	Valence Electrons	3	4	5	6	7	8
1	1	1	2	Family	3	4	5	6	7	8
1	1	1	2	Family	3	4	5	6	7	8
3	4	5	6	7	8	9	10	11	12	13
Li	Be	B	C	N	O	F	Ne	Na	Mg	Al
Lithium	Beryllium	Boron	Carbon	Nitrogen	Oxygen	Fluorine	Neon	Sodium	Magnesium	Aluminum
6.941	9.0122	10.81	12.011	14.007	15.999	18.998	20.180	22.990	24.305	26.981
11	12	13	14	15	16	17	18	19	20	21
Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca	Sc
Sodium	Magnesium	Aluminum	Silicon	Phosphorus	Sulfur	Chlorine	Argon	Potassium	Calcium	Scandium
22.990	24.305	26.981	28.086	30.974	32.06	35.453	39.948	39.098	40.078	44.956
19	20	21	22	23	24	25	26	27	28	29
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu
Potassium	Calcium	Scandium	Titanium	Vanadium	Chromium	Manganese	Iron	Cobalt	Nickel	Copper
39.098	40.078	44.956	47.88	50.942	51.996	54.938	55.845	58.933	58.693	63.546

8. The oxidation number of an ion is the number of electrons an atom will lose to
become a **negative positive** ion or gain to become a **negative positive** ion and
have a full outer energy level of electrons.

Combine the elements or polyatomic compounds and write the formula and the name

	Symbols & Oxidation #	Formula	Name of Compound
9. Sodium and phosphate			
10. Magnesium and oxygen			
11. Potassium and phosphorus			
12. Ammonium and hydroxide			
13. Sodium and chlorine			
14. Beryllium and bromine			