

1+						0								
1						8								
1	2+	Oxidation Number	3+	4+/-	3-	2-	1-	8						
1 H Hydrogen 1.008	2- 2	Valence Electrons Family	3	4	5	6	7	4 He Helium 4.003						
3 Li Lithium 6.941	4 Be Beryllium 9.0122	<div>Some Polyatomic Ions and their Oxidation Numbers</div> <div><div>1+ ammonium (NH₄)</div><div><div>1- acetate (C₂H₃O₂)</div><div>chlorate (ClO₃)</div><div>hydroxide (OH)</div><div>nitrate (NO₃)</div><div>bicarbonate(HNO₃)</div></div><div><div>2- carbonate (CO₃)</div><div>sulfate (SO₄)</div></div><div><div>3- phosphate (PO₄)</div></div></div>							5 B Boron 10.81	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.305								13 Al Aluminum 26.98	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.06	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078								31 Ga Gallium 69.72	32 Ge Germanium 72.61	33 As Arsenic 74.922	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80

	Symbols & Oxidation #	Formula	Name of Compound
19. Sodium and chlorine			
20. Magnesium and oxygen			
21. Potassium and chlorate			
22. Ammonium and hydroxide			
23. Potassium and phosphorus			
24. Aluminum and bromine			

25. An ion is an atom or group of atoms that has become electrically _____

26. When an atom loses an electron its charge is (**positive or negative**)

27. An ionic bond is the attraction between (**opposites, positive, neutral, or negative**) ions.

28. Ionic compounds are electrically (**charged, positive, neutral, or negative**).

29. The sum of the charges for an ionic compound is _____.

The two answers must be in the right order.

30. An ionic compound is the result of the bonding of a (**non-metal, metalloid, metal, noble gas**) with a (**non-metal, metalloid, metal**).