

**BOYLE'S LAW**

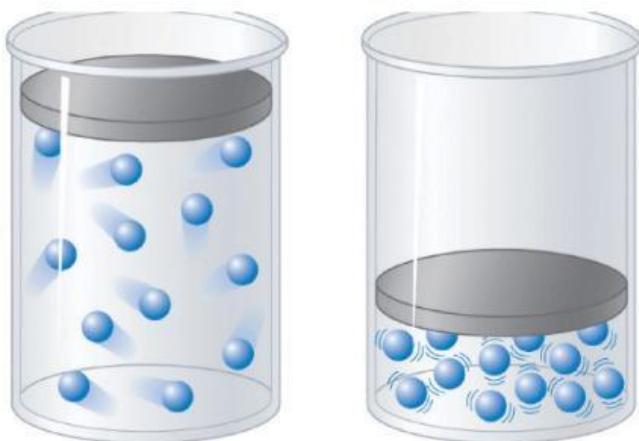
Total questions: 10

Worksheet time: 22mins

Instructor name: Khaled Barhoom

1. What is the formula for Boyle's Law?  
a)  $P_1V_1=P_2V_2$       b)  $P_1V_1/P_2V_2$   
c)  $P_1V_2=P_2V_1$       d)  $P_1/V_1=P_2/V_2$
  
2. When Pressure increases then the Volume must...  
a) Increase      b) decrease
  
3. When Volume increases then Pressure must...  
a) Increase      b) Decrease
  
4. A gas at a volume of 4 liters is at a pressure of 2 atm. The volume is changed to 16 Liters, what must the new pressure be?  
a) 2 atm      b) 12 atm  
c) 10 atm      d) 0.5 atm

5.



Which container will have a lower pressure?

a) left      b) right  
c) they both have the same pressure      d) I don't know

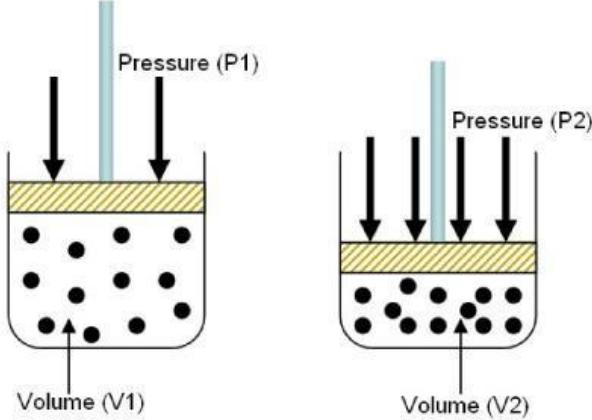
6. What type of relationship to pressure and volume have?

- a) direct
- b) inverse
- c) no relationship
- d) I don't know

7. The relationship of which two variables are compared in Boyle's Law?

- a) pressure & volume
- b) volume & temperature
- c) temperature & pressure
- d) volume & moles (amount of gas)

8.



What happened to the volume between the two steps?

- a) volume increased
- b) volume decreased
- c) volume stayed the same

9. True or False: Gases can be compressed.

- a) True
- b) False

10. The relationship of which two variables are compared in Boyle's Law?

- a) pressure & volume
- b) volume & temperature
- c) temperature & pressure
- d) volume & moles (amount of gas)