___ Cu + ___ S_8 → ___ CuS

Type of Reaction:_____

___ Al_2O_3 → ___Al + ___O_2

Type of reaction: _____

___ $AlCl_3$ + ___ Mg → ___ $MgCl_2$ + ___Al

Type of Reaction:_____

___ Na_3N + ___$CaCl_2$ → ____ NaCl + ___ Ca_3N_2

Type of reaction: _____

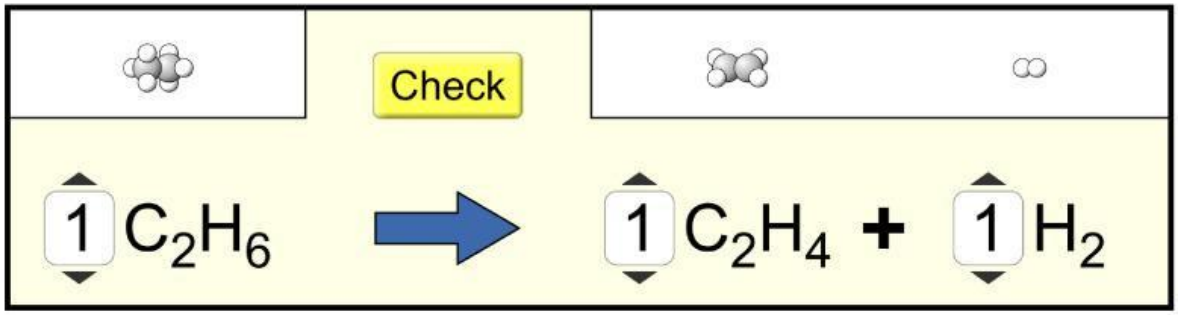
___C_3H_8 + ___O_2 → ___CO_2 + ___H_2O + Heat

Type of Reaction:_____

___ $Al(ClO_3)_3$ → ____ $AlCl_3$ + ___O_2

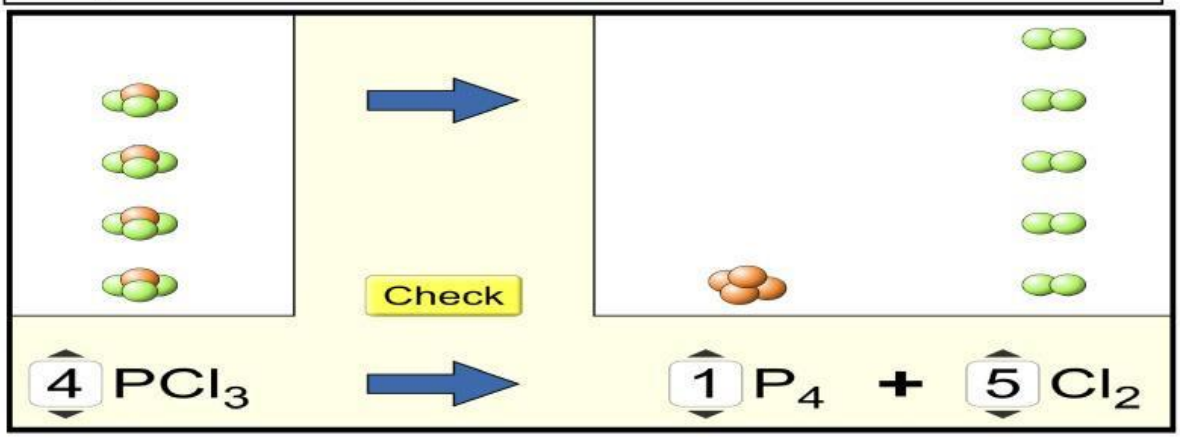
Type of reaction: _____

LIVEWORKSHEETS



How many of each element do we have on the reactants side of the equation?		How many of each element do we have on the products side of the equation?	
___ C atoms	(colored _____)	___ C atoms	(colored _____)
___ H atoms	(colored _____)	___ H atoms	(colored _____)

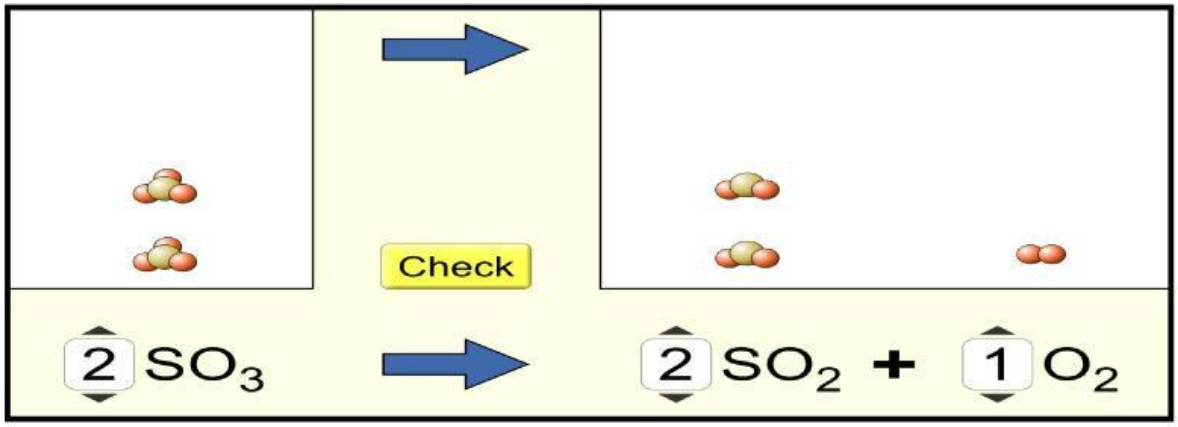
Because _____ have the same number of each element on each side, we say that the chemical equation above _____ Balanced.



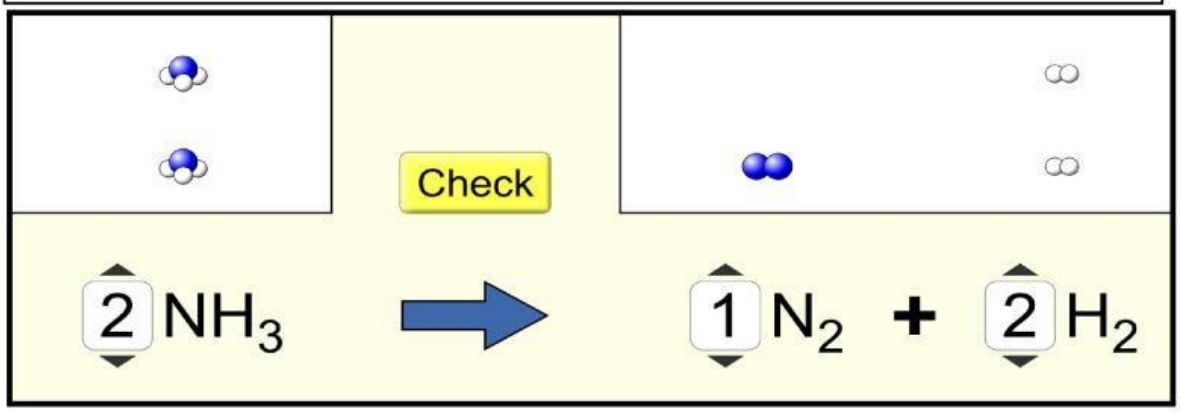
How many of each element do we have on the reactants side of the equation		How many of each element do we have on the products side of the equation	
___ P atoms	(colored _____)	___ P atoms	(colored _____)
___ Cl atoms	(colored _____)	___ Cl atoms	(colored _____)

Because _____ have the same number of each element on each side, we say that the chemical equation above _____ Balanced

LIVEWORKSHEETS



How many of each element do we have on the reactants side of the equation	How many of each element do we have on the products side of the equation
___ S atoms (colored _____)	___ S atoms (colored _____)
___ O atoms (colored _____)	___ O atoms (colored _____)
Because _____ have the same number of each element on each side, we say that the chemical equation above _____ Balanced	



How many of each element do we have on the reactants side of the equation	How many of each element do we have on the products side of the equation
___ N atoms (colored _____)	___ N atoms (colored _____)
___ H atoms (colored _____)	___ H atoms (colored _____)
Because _____ have the same number of each element on each side, we say that the chemical equation above _____ Balanced	

LIVEWORKSHEETS