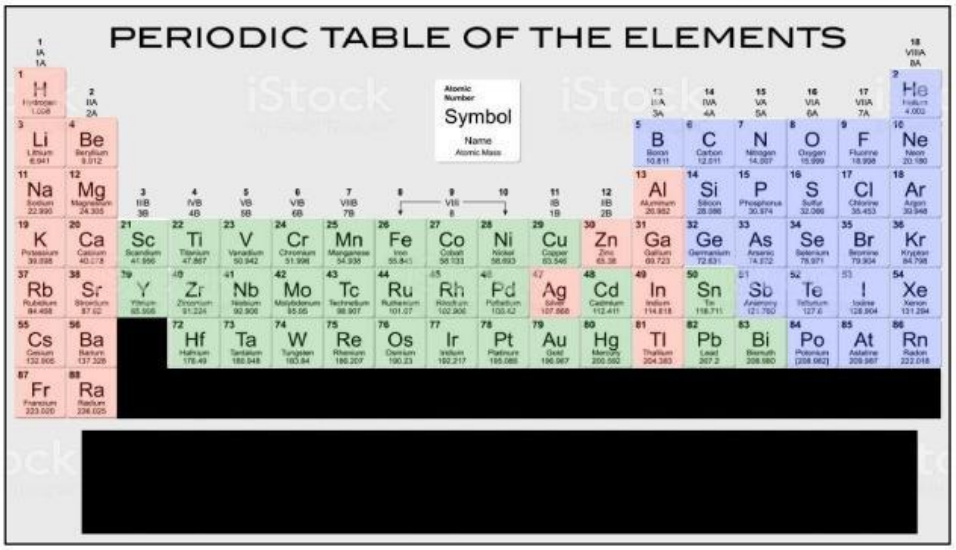
Periodic Table with bonding rules based on location of first element

+1 | +2 | |+3| N/A |-3 |-2 | -1 | 0



Polyatomic Ions

Polyatomic Ions	
NH_4^+	Ammonium
BrO_3^-	Bromate
CN^-	Cyanide
$C_2H_3O_2^-$ (CH_3COO^-)	Acetate
ClO_4^-	Perchlorate
ClO_3^-	Chlorate
ClO_2^-	Chlorite
ClO^-	Hypochlorite
IO_3^-	Iodate
MnO_4^-	Permanganate
NO_3^-	Nitrate
NO_2^-	Nitrite
OH^-	Hydroxide
HCO_3^-	Hydrogen carbonate
HSO_4^-	Hydrogen sulfate
SCN^-	Thiocyanate
CO_3^{2-}	Carbonate
$Cr_2O_7^{2-}$	Dichromate
CrO_4^{2-}	Chromate
SO_4^{2-}	Sulfate
SO_3^{2-}	Sulfite
PO_4^{3-}	Phosphate

NOTE THE DIFFERENCES BETWEEN THE ENDINGS

Cl = Chlorine
Cl^-1 = Chloride
ClO3^-1 = Chlorate
ClO2^-1 = Chlorite

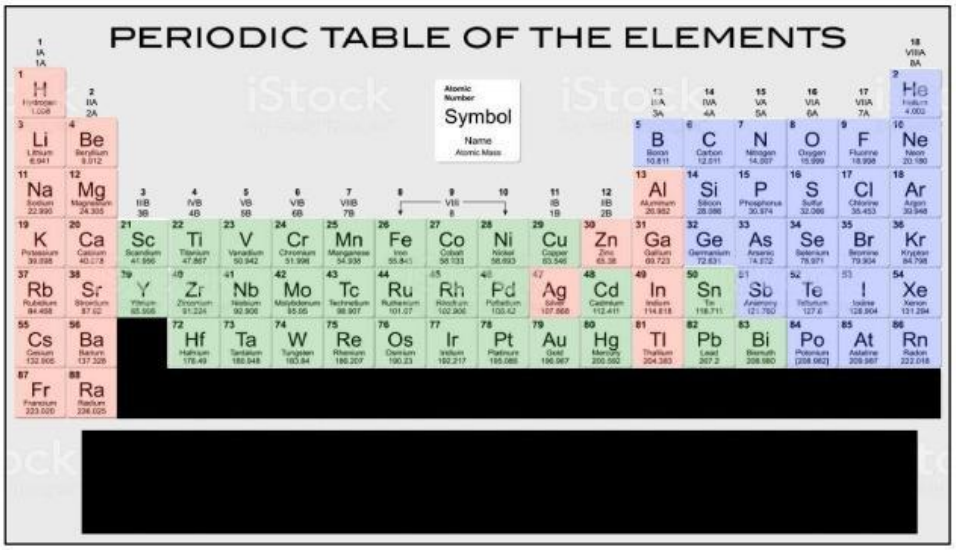
- Silver (Ag) always has a charge of _____ and Zinc (Zn) always has a charge of _____.
- If the first element in the compound is red, you _____need a roman numeral in the chemical name.
- If the first element in the compound is green, you _____need a roman numeral in the chemical name.

Roman Numeral needed?	Name	Ions	Formula
		Mg^{+2} ClO_3^-	
	Zinc Chloride		
			$Ag_2Cr_2O_7$
			MnN
	Iron(III) Nitrate		
			Sn_2O_3

LIVEWORKSHEETS

Periodic Table with bonding rules based on location of first element

+1 | +2 | |+3| N/A |-3 |-2 | -1 | 0



Polyatomic Ions

Polyatomic Ions	
NH_4^+	Ammonium
BrO_3^-	Bromate
CN^-	Cyanide
$C_2H_3O_2^-$ (CH_3COO^-)	Acetate
ClO_4^-	Perchlorate
ClO_3^-	Chlorate
ClO_2^-	Chlorite
ClO^-	Hypochlorite
IO_3^-	Iodate
MnO_4^-	Permanganate
NO_3^-	Nitrate
NO_2^-	Nitrite
OH^-	Hydroxide
HCO_3^-	Hydrogen carbonate
HSO_4^-	Hydrogen sulfate
SCN^-	Thiocyanate
CO_3^{2-}	Carbonate
$Cr_2O_7^{2-}$	Dichromate
CrO_4^{2-}	Chromate
SO_4^{2-}	Sulfate
SO_3^{2-}	Sulfite
PO_4^{3-}	Phosphate

NOTE THE DIFFERENCES BETWEEN THE ENDINGS

Cl = Chlorine
Cl^-1 = Chloride
ClO3^-1 = Chlorate
ClO2^-1 = Chlorite

Roman Numeral Needed?	Name	Ions	Formula
		Al^{+3} MnO_4^{-1}	
	Lithium Sulfide		
		Co^{+3} Br^{-1}	
			$MnCO_3$
	Vanadium(IV) Sulfite		
			$TiBr_2$
		Zn^{+2} ClO^{-1}	
	Calcium Bromate		

LIVEWORKSHEETS

Metals Review:
Metals are located on the _____ of the periodic table.

Metals have _____ electronegativities which means they are _____ at taking other atoms' electrons.

Metals have _____ ionization energies which means it is _____ to take their electrons.

Metals tend to _____ electrons to become _____ charged _____.

Nonmetals Review:
Nonmetals are located on the _____ of the periodic table.

Nonmetals have _____ electronegativities which means they are _____ at taking other atoms' electrons.

Nonmetals have _____ ionization energies which means it is _____ to take their electrons.

Nonmetals tend to_____ electrons to become_____charged _____.

LIVEWORKSHEETS