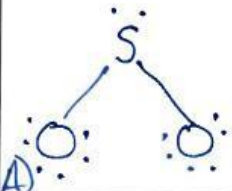
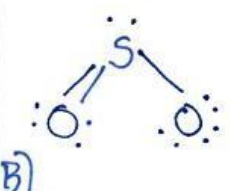
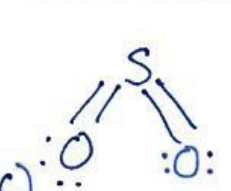
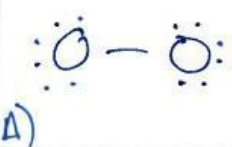
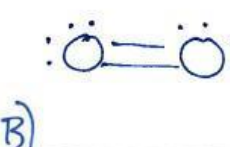
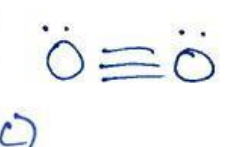
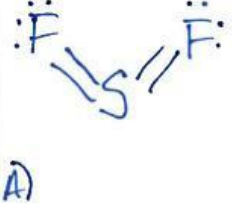
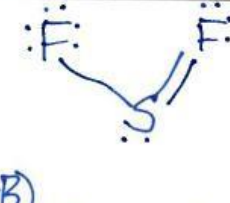
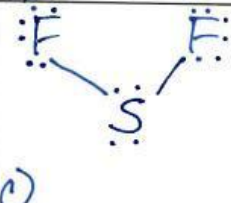
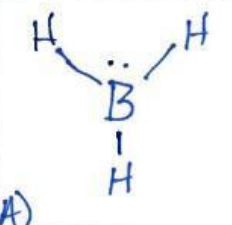
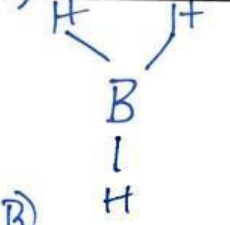
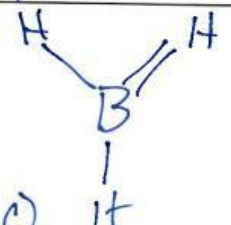


Choose the proper Lewis Structure for each:

nitrogen trifluoride, NF ₃	A)	B)	C)
hydrogen sulfide, H ₂ S	A)	B)	C)
fluorine, F ₂	A)	B)	C)
carbon monoxide, CO	A)	B)	C)

Choose the proper Lewis Structure for each:

sulfur dioxide, SO ₂	 A)	 B)	 C)
oxygen, O ₂	 A)	 B)	 C)
sulfur difluoride, SF ₂	 A)	 B)	 C)
boron trihydride, BH ₃	 A)	 B)	 C)

Choose the proper Lewis Structure for each:

chloroform, CHCl_3	$\begin{array}{c} \text{H} \\ \\ \text{Cl} = \text{C} = \text{Cl} \\ \\ \text{Cl} \end{array}$ A)	$\begin{array}{c} \text{H} \\ \\ \text{Cl} - \text{C} - \text{Cl} \\ \\ \text{Cl} \end{array}$ B)	$\begin{array}{c} \text{H} \\ \\ \text{Cl} - \text{C} - \text{Cl} \\ \\ \text{Cl} \end{array}$ C)
carbon disulfide, CS_2	$\begin{array}{c} \cdot\cdot \\ \text{S} = \text{C} = \text{S} \\ \cdot\cdot \end{array}$ A)	$\begin{array}{c} \cdot\cdot \\ \text{S} - \text{C} - \text{S} \\ \cdot\cdot \end{array}$ B)	$\begin{array}{c} \cdot\cdot \\ \text{S} - \text{C} = \text{S} \\ \cdot\cdot \end{array}$ C)
beryllium chloride, BeCl_2	$\begin{array}{c} \cdot\cdot \\ \text{Cl} - \text{Be} - \text{Cl} \\ \cdot\cdot \end{array}$ A)	$\begin{array}{c} \cdot\cdot \\ \text{Cl} - \text{Be} = \text{Cl} \\ \cdot\cdot \end{array}$ B)	$\begin{array}{c} \cdot\cdot \\ \text{Cl} - \text{Be} - \text{Cl} \\ \cdot\cdot \end{array}$ C)
hydrogen cyanide, HCN	$\text{H} - \text{C} \equiv \text{N}$ A)	$\text{H} = \text{C} = \text{N}$ B)	$\text{H} \equiv \text{C} - \text{N}$ C)

Choose the proper Lewis Structure for each:

acetylene, C ₂ H ₂	$\text{H}-\ddot{\text{C}}=\ddot{\text{C}}-\text{H}$ A)	$\text{H}-\text{C}\equiv\text{C}-\text{H}$ B)	$\text{H}-\ddot{\text{C}}-\ddot{\text{C}}-\text{H}$ C)
silicon dioxide, SiO ₂	$\begin{array}{c} \ddot{\text{O}} \\ \vdots \\ \text{O}=\ddot{\text{Si}}=\text{O} \\ \vdots \\ \ddot{\text{O}} \end{array}$ A)	$\begin{array}{c} \ddot{\text{O}} \\ \vdots \\ \text{O}=\ddot{\text{Si}}=\text{O} \\ \vdots \\ \ddot{\text{O}} \end{array}$ B)	$\begin{array}{c} \ddot{\text{O}} \\ \vdots \\ \text{O}=\ddot{\text{Si}}=\text{O} \\ \vdots \\ \ddot{\text{O}} \end{array}$ C)
hydrogen peroxide, H ₂ O ₂	$\text{H}-\text{O}=\text{O}-\text{H}$ A)	$\text{H}-\ddot{\text{O}}=\ddot{\text{O}}-\text{H}$ B)	$\text{H}-\ddot{\text{O}}-\ddot{\text{O}}-\text{H}$ C)
sulfate, SO ₄ ²⁻	$\left[\begin{array}{c} \ddot{\text{O}} \\ \vdots \\ \text{O}=\ddot{\text{S}}=\text{O} \\ \vdots \\ \ddot{\text{O}} \end{array} \right]^{2-}$ A)	$\left[\begin{array}{c} \ddot{\text{O}} \\ \vdots \\ \text{O}=\ddot{\text{S}}=\text{O} \\ \vdots \\ \ddot{\text{O}} \end{array} \right]^{2-}$ B)	$\begin{array}{c} \ddot{\text{O}} \\ \vdots \\ \text{O}=\ddot{\text{S}}=\text{O} \\ \vdots \\ \ddot{\text{O}} \end{array}$ C)

Choose the proper Lewis Structure for each:

methanol, CH ₃ OH	A)	B)	C)
nitrate, NO ₃ ⁻	A)	B)	C)
chlorite, ClO ₂ ⁻	A)	B)	C)
formic acid, CH ₂ O	A)	B)	C)