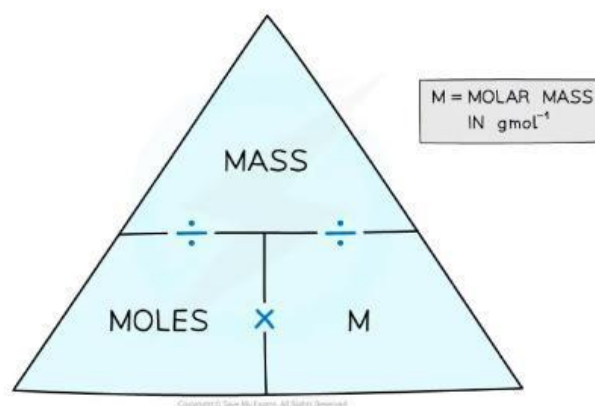


### Part 1: Counting Atoms

H <sub>2</sub> O	# of atoms	NaCl	# of atoms	O <sub>2</sub>	# of atoms	CaCO <sub>3</sub>	# of atoms	C <sub>6</sub> H <sub>12</sub> O <sub>6</sub>	# of atoms
H		Na		O		Ca		C	
O		Cl				C		H	
						O		O	

**Part 2: Molar Mass:** Use the information above and your new periodic table to calculate the molar mass of all five compounds.

1. H<sub>2</sub>O g/mol
2. NaCl g/mol
3. O<sub>2</sub> g/mol
4. CaCO<sub>3</sub> g/mol
5. C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> g/mol



**Part 3: Mole Mass Conversions:** Use your mole to mass triangle (see above) and molar masses to solve the following:

1. Convert 5.00 moles of H<sub>2</sub>O to grams.
2. Convert 20.0 grams of NaCl to moles.
3. Convert 3.50 moles of O<sub>2</sub> to grams.
4. Convert 15.0 grams of CaCO<sub>3</sub> to moles.
5. Convert 2.00 moles of C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> to grams.