

Chemical Equation

Name :

Class :

Balance the chemical equation below!

1	$\dots \text{H}_2 + \dots \text{Br}_2 \longrightarrow \dots \text{HBr}$
2	$\dots \text{N}_2 + \dots \text{H}_2 \longrightarrow \dots \text{NH}_3$
3	$\dots \text{K} + \dots \text{H}_2\text{O} \longrightarrow \dots \text{KOH}$
4	$\dots \text{S}_8 + \dots \text{O}_2 \longrightarrow \dots \text{SO}_2$
5	$\dots \text{H}_2\text{O}_2 \longrightarrow \dots \text{H}_2\text{O} + \dots \text{O}_2$
6	$\dots \text{P}_4\text{O}_{10} + \dots \text{H}_2\text{O} \longrightarrow \dots \text{H}_3\text{PO}_4$
7	$\dots \text{Zn} + \dots \text{AgCl} \longrightarrow \dots \text{ZnCl}_2 + \dots \text{Ag}$
8	$\dots \text{Al} + \dots \text{H}_2\text{SO}_4 \longrightarrow \dots \text{Al}_2(\text{SO}_4)_3 + \dots \text{H}_2$
9	$\dots \text{NaHCO}_3 \longrightarrow \dots \text{Na}_2\text{CO}_3 + \dots \text{H}_2\text{O} + \dots \text{CO}_2$
10	$\dots \text{CaCO}_3 + \dots \text{HCl} \longrightarrow \dots \text{CaCl}_2 + \dots \text{H}_2\text{O} + \dots \text{CO}_2$