

# STRONG ACID

The concentration of HCl is 0.2 M.

Calculate the pH of solution

$$\begin{aligned}\text{pH} &= -\log [\text{H}^+] \\ &= -\log \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}}\end{aligned}$$

## STRONG ACID

The concentration of  $\text{H}_2\text{SO}_4$  is 0.2 M.

Calculate the pH of solution

$$\begin{aligned}\text{pH} &= -\log [\text{H}^+] \\ &= -\log \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}}\end{aligned}$$