Limiting Reactant Stoichiometry

Q#1: Plants use carbon dioxide and water vapors to generate glucose by photosynthesis. If a plant absorbs 440 grams of CO_2 and 360 grams of H_2O in one hour, then how much glucose plant will produce?

$6CO_2 + 6H_2O \rightarrow C_6H_{12}O_6 + 6O_2$		
1. Moles of reactants	$CO_2 = \frac{}{44} =$	$H_2O = \frac{}{18} =$
2. Divide moles on coefficient of reactant.	$CO_2 = - =$	$H_2O = -$
3. Limiting Reactant		
4. Moles of glucose	$- \times \quad =$	
5. Mass of glucose		

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