Identification of Limiting Reactant

Q#1: The reaction between NaOH and H_2SO_4 produces Na_2SO_4 and H_2O as products. If 4 grams of NaOH reacts with 9.8 grams of H_2SO_4 then find the limiting reactant.

$2NaOH + H_2SO_4 \rightarrow Na_2SO_4 + 2H_2O$		
1. Moles of reactants	$NaOH = \frac{}{40} =$	$H_2SO_4 = \frac{}{98} =$
2. Divide moles on coefficient of reactant.	$NaOH = \frac{}{2}$	$H_2SO_4 = \frac{}{1}$
3. Limiting Reactant		

Q#2: The reaction between Hydrogen and Oxygen produces water.

a) **Balance the Chemical Equation for the reaction.**
b) **If 0.8 grams of H_2 reacts with 1.6 grams of O_2 then find the limiting reactant.**

$H_2 + O_2 \rightarrow H_2O$		
1. Moles of reactants	$Hydrogen = \frac{}{2} =$	$Oxygen = \frac{}{32} =$
2. Divide moles on coefficient of reactant.	$Hydrogen = -$	$Oxygen = — =$
3. Limiting Reactant		