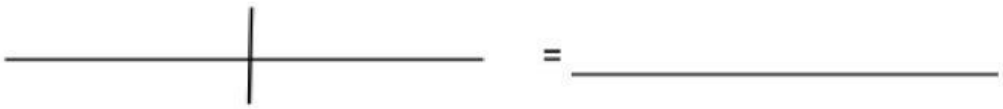
Name _____
Period ___ Date ___/___/___

Mixed Mole Practice

1 mole = 6.02 x 10^{23} particles
1 mole = molar mass

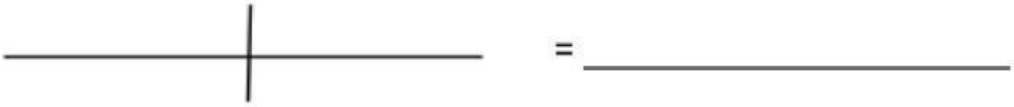
1. Calculate the mass of 1.58 moles CH_4. [molar mass CH_4 = 16.05 g/mol]
   Given: 1.58 moles CH_4
   Unknown: ? g CH_4
   Will you use molar mass or 6.02 x 10^{23} to solve this problem?



3. How many molecules are there in a 0.583 mole sample of H_2O? [molar mass of H_2O = 18.02 g/mol]
   G: 0.583 moles H_2O
   U: ? molecules H_2O
   Will you use molar mass or 6.02 x 10^{23} to solve this problem?



4. How many moles of 5.79 x 10^{20} molecules of CO_2 ? [molar mass CO_2 = 44.01 g/mol]
   G: 5.79 x 10^{20} molecules CO_2
   U: ? mole CO_2
   Will you use molar mass or 6.02 x 10^{23} to solve this problem?

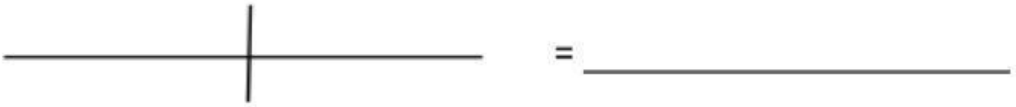


5. How many moles are in a 35.0 gram sample of H_2O? [molar mass H_2O = 18.02 g/mol]
   G: 35.0 g H_2O
   U: ? moles H_2O
   Will you use molar mass or 6.02 x 10^{23} to solve this problem?



LIVEWORKSHEETS

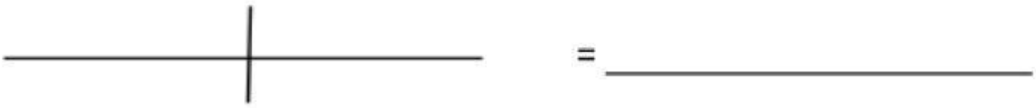
6. How many grams of NaOH do you measure if you need 2.87 moles of NaOH?



7. How many moles of NaCl are in 2.11×10^{24} particles of NaCl?



8. How many molecules are present in 1.45 moles of H_2O?



9. If you have 10.33 grams of copper, how many moles of copper is that?



10. Calculate the molar mass of $Al(C_2H_3O_2)_3$.

LIVEWORKSHEETS