Ionic Bonding

Binary Ionic Nomenclature

Ionic bonds are bonds that involve a **METAL** and a **NONMETAL.**

When writing the chemical formulas for ionic bonds we must:

1. Find the oxidation number/charge for each element
2. The metal/cation always comes first
3. Cross the oxidation numbers/charges so that they are now subscripts on the other element.
    a. Do not record 1 as a subscript.
    b. Reduce subscripts if possible.

*For the chart below write the **chemical formula** for each ionic bond.

	C	N	P	S	F
Li	Li_4C				
Ca					
K					
Sr					
Al					

Ionic Bonding

When naming ionic bonds we must first write the name of the metal and then the name of the nonmetal. The ending of the nonmetal needs to be changed to -ide.

Example: K and Se are bonded together. We would call this **potassium selenide**.
* In the chart below write the names of each ionic bond.

	C	N	S	Cl	F
Li	lithium carbide				
Ca					
K					
Sr					
Al					