Worksheet - Titration Problems

1. What is the M of NaOH if it takes 40 ml of NaOH to reach the equivalence point in a titration with 50 ml of 0.2 M HCl?

M

2. 50 ml of 0.3 M KOH are required to titrate 60 ml of H_2SO_4. What is the M of the H_2SO_4?

M

3. 60 ml of 1.2 M NaOH are required to titrate 40 ml of HF. What is the M of the HF?

M

4. What volume of 0.40 M NaOH would be required to titrate 100 ml of 0.25 M HCl?

ml

5. 40 ml of 0.1M H_3PO_4 are required to titrate 150 ml of NaOH to the equivalence point. What is the M of the NaOH?

M

6. 55 ml of 1.2 M H $C_2H_3CO_2$ are used to titrate a sample of 0.67 M $Ba(OH)_2$. What is volume of the $Ba(OH)_2$ used?

ml

7. 90 ml of 0.25 M $Ca(OH)_2$ are required to titrate 100 ml of HCl. What is M of the HCl?

M

8. 50 ml of 0.45M $Sr(OH)_2$ are required to titrate a .75 M H_2SO_4 sample. What is the volume of the H_2SO_4?

ml

9. A 94 mL sample of citric acid, $H_3C_6H_5O7$, solution (example=orange juice) is titrated to the phenolphthalein endpoint using 7.0 mL of 0.010 M NaOH. What is the concentration of the citric acid in the orange juice?

M

10. 30 ml of 0.3 M NaOH are required to titrate H_3PO_4 to the equivalence point. How many moles of H_3PO_4 are needed to reach the equivalence point?

moles H_3PO_4

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