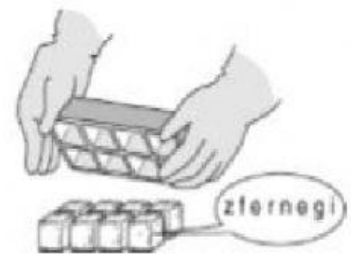
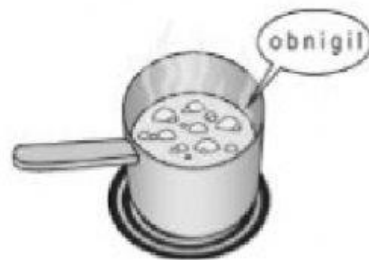
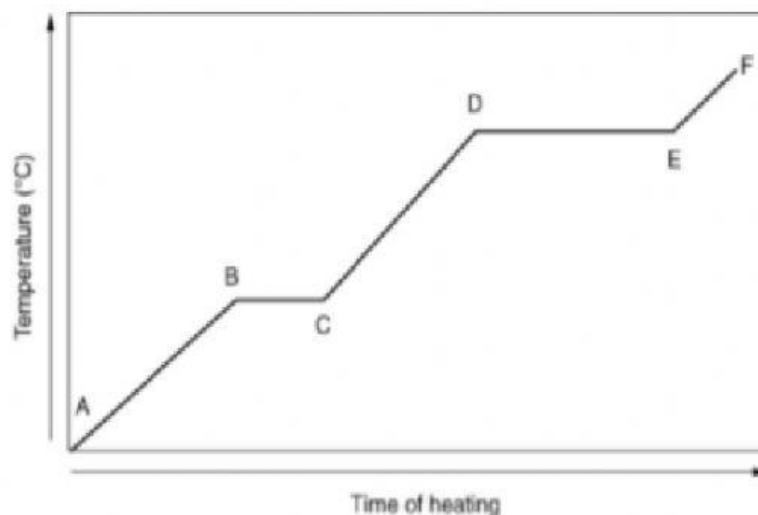


Arrange the jumbled letters to make a word that describes the change shown in the drawings.



Look at the graph shown below, and then answer the questions which follow.



- What is happening to the temperature in section E to F?
- What is the state of the substance in section A to B?
- What change of state is occurring in section D to E?
- Explain why the temperature doesn't rise in section B to C although heating continues.

The melting points of elements in the periodic table are shown below.

H -259.1						
Li 180.5						He -268.93
Na 97.8						
K 63.5	Ca 842					
Rb 39.3	Sr 777					
Cs 28.44	Ba 727					
Fr 27	Ra 900					

B 2075	C 3629 ap	N -210	O -218.79	F -219.68	Ne -248.603
Al 960.32	Si 1414	P 44.15	S 115.21	Cl -101.5	Ar -189.36
Ga 29.76	Ge 937	As 817	Se 221	Br -157.26	Kr -157.26
In 156.6	Sn 231.93	Sb 630.63	Te 449.51	I 113.7	Xe -111.74
Tl 304	Pb 327.46	Bi 271.4	Po 254	At 302	Rn -71

- Name four elements in the halogen group.
- Describe the general trend in melting points down the alkali metal and noble gas groups.
- What is the general trend in melting points down group 6 from oxygen to polonium? Identify any melting point within this group that does not seem to fit the trend.
- The melting points of germanium and bromine are missing. Estimate these melting points and explain your reasoning.

Tick (✓) the correct boxes to show the typical properties of metals and non-metals.

	Good conductors of electricity	Most are brittle when solid	Most have high melting points
Metals			
Non-metals			