

Le Chatelier's Equilibrium Practice

Fill out the chart using the following equation based on the rules of Le Chatelier's Principle:

*The ΔH is _____ so this is an _____ reaction, so I will write the word HEAT on the _____ with the _____.

*There are more moles of gas on the _____ side of the arrow so I will write the word PRESSURE on the _____ with the _____ and I will write the word VOLUME on the opposite side.



Stress	Equilibrium Shift L or R	$[\text{C}_3\text{H}_8]$ increase or decrease	$[\text{O}_2]$ increase or decrease	$[\text{CO}_2]$ increase or decrease	$[\text{H}_2\text{O}]$ increase or decrease	Temp/Heat increase or decrease
Add O_2						
$30^\circ\text{C} \rightarrow 15^\circ\text{C}$						
Remove C_3H_8						
Increase pressure						
Increase carbon dioxide						
$4\text{L} \rightarrow 8\text{L}$						
Remove water						
Heat added						
Decrease pressure						