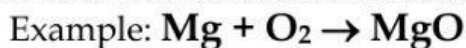


# REDOX IN EQUATIONS

Determine what is oxidized and what is reduced in the following equations, using oxidation states. Complete the statements that follow each equation.

An example is provided for you below:



**Mg** is *oxidized* because its oxidation state *increased* from 0 in **Mg** to +2 in **MgO**.

**O<sub>2</sub>** is *reduced* because its oxidation state *decreased* from 0 in **O<sub>2</sub>** to -2 in **MgO**.



\_\_\_\_\_ is *oxidized* because its oxidation state *increased* from \_\_\_\_\_ in \_\_\_\_\_

to \_\_\_\_\_ in \_\_\_\_\_.

\_\_\_\_\_ is *reduced* because its oxidation state *decreased* from \_\_\_\_\_ in \_\_\_\_\_

to \_\_\_\_\_ in \_\_\_\_\_.

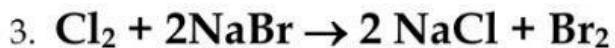


\_\_\_\_\_ is *oxidized* because its oxidation state *increased* from \_\_\_\_\_ in \_\_\_\_\_

to \_\_\_\_\_ in \_\_\_\_\_.

\_\_\_\_\_ is *reduced* because its oxidation state *decreased* from \_\_\_\_\_ in \_\_\_\_\_

to \_\_\_\_\_ in \_\_\_\_\_.



\_\_\_\_\_ is *oxidized* because its oxidation state *increased* from \_\_\_\_\_ in \_\_\_\_\_  
to \_\_\_\_\_ in \_\_\_\_\_.

\_\_\_\_\_ is *reduced* because its oxidation state *decreased* from \_\_\_\_\_ in \_\_\_\_\_  
to \_\_\_\_\_ in \_\_\_\_\_.



\_\_\_\_\_ is *oxidized* because its oxidation state *increased* from \_\_\_\_\_ in \_\_\_\_\_  
to \_\_\_\_\_ in \_\_\_\_\_.

\_\_\_\_\_ is *reduced* because its oxidation state *decreased* from \_\_\_\_\_ in \_\_\_\_\_  
to \_\_\_\_\_ in \_\_\_\_\_.



\_\_\_\_\_ is *oxidized* because its oxidation state *increased* from \_\_\_\_\_ in \_\_\_\_\_  
to \_\_\_\_\_ in \_\_\_\_\_.

\_\_\_\_\_ is *reduced* because its oxidation state *decreased* from \_\_\_\_\_ in \_\_\_\_\_  
to \_\_\_\_\_ in \_\_\_\_\_.