

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

Number of Zn atoms = **number mole** × *Avogadros number*

Number of Zn atoms = × 6.02×10^{23}

Answer = $\times 10^{24}$

2. Calculate the number of molecules in 11.5 mol of water (H₂O).

Number of H₂O molecules = **number mole** × *Avogadros number*

Number of H₂O molecules = × $\times 10^{23}$

Answer = $\times 10^{24}$