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1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

Number of Zn atoms = number mole × Avogadros number

Number of Zn atoms = \times 6.02x10²³

Answer = $x10^{24}$

2. Calculate the number of molecules in 11.5 mol of water (H₂O).

Number of H₂O molecules = number mole × Avogadros number

Number of H_2O molecules = \times $\times 10^{23}$

Answer = $x10^{24}$