

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

Number of Zn atoms =  $\text{number mole} \times \text{Avogadros number}$

Number of Zn atoms =  $\quad \times 6.02 \times 10^{23}$

Answer =  $\quad \times 10^{24}$

2. Calculate the number of molecules in 11.5 mol of water ( $\text{H}_2\text{O}$ ).

Number of  $\text{H}_2\text{O}$  molecules =  $\text{number mole} \times \text{Avogadros number}$

Number of  $\text{H}_2\text{O}$  molecules =  $\quad \times \quad \times 10^{23}$

Answer =  $\quad \times 10^{24}$